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GRADIENT ANALYSIS OF CARBON MONOXIDE AND
METHANE IN POLLUTED AND OTHER
NEARSHORE HABITATS

James Taylor Welch

Gradient Analysis of Carbon Monoxide
and Methane in Polluted and Other
Nearshore Habitats

by

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Lieutenant, United States Navy
B.S.Che., Purdue University, 1966

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ABSTRACT

A system for the determination of dissolved gases in seawater by gas chromatography was constructed and used to find the concentrations of methane and carbon monoxide in a variety of habitats around the Monterey Peninsula. Methane was shown to have a maximum of 2.8×10^{-4} ml/l at 50 meters at the open ocean station, with a surface value of 1.1×10^{-4} ml/l. The surface waters at the nearshore stations were almost three times this value. Methane was also shown to be an effective tracer for sewage effluent. The carbon monoxide maximum of 2.1×10^{-4} ml/l was found at 15 meters which correlated closely with primary productivity (Rowney 1973). The surface value of 0.81×10^{-4} ml/l was lower than the nearshore values. All stations sampled were found to be highly supersaturated with both gases. This indicates that in this area, the ocean is a major source of both methane and carbon monoxide.

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I. INTRODUCTION

The ocean has been shown to be supersaturated with carbon monoxide (Swinnerton, Linnenbom, and Lamontagne, 1970). The sources for this gas may be plants (Chapman and Tocher, 1966; Delwiche, 1970; Loewus and Delwiche, 1963), animals (Pickwell and Barham, 1964; Pickwell, 1970; Barham, 1963; Wittenberg, 1960), and microorganisms (Junge, et. al., 1971). These results led to the hypothesis that carbon monoxide production might be related to primary productivity. The highly productive waters of Monterey Bay were thought to be an excellent location to test this hypothesis.

Methane has also been reported as being present in surface waters (Swinnerton, Linnenbom, and Cheek, 1969). Since this gas is a product of anaerobic decomposition of organic matter it was felt that it might be useful as a pollution tracer from sewage outfalls. Again, Monterey Bay provides an excellent environment for these studies.

In order to measure these gases, the highly sensitive methods of gas chromatography were used. The gas chromatograph has long been one of the analytical chemist's most useful instruments. It was not until 1962 that a practical system for oceanographic analyses was developed (Swinnerton, Linnenbom, and Cheek, 1962). Today, the shipboard determination of dissolved gases by gas chromatography is one of the most valuable methods available to the chemical oceanographer.

II. EQUIPMENT

The system used was essentially that described by Swinnerton, Linnenbom, and Cheek (1968). A calibration system was added and minor modifications were made in the sample transfer system. The entire system is shown in Figure 1.

A. GAS TRAPPING SYSTEM

The separation is accomplished in four major steps. They are sample transfer, stripping, trapping, and backflushing.

1. Sample Transfer

A helium line is connected to the side port of a Swagelok heat exchanger "T". A standard taper joint is attached to the bottom of this "T" that fits into the filled sample bottle. A 1/8" stainless steel tube runs from the bottom of the sample bottle through the heat exchanger "T" and is connected to the bottom of the stripping chamber.

Transfer of the sample is accomplished by displacement with helium. The helium flow is controlled by a micrometer valve. The helium forces the water sample through the stainless steel tube into the stripping chamber. A toggle valve is located in the helium line to release the pressure in the sample bottle after transfer.

2. Stripping Chamber

The stripping chamber is a glass tube 65 mm. in diameter and 45 cm. tall. It is fitted with a coarse fritted disc at the bottom and a 10 cm. neck, 2 cm. in diameter, at the top. "Purge helium" is introduced just below the fritted disc. It is controlled by a Teflon stopcock and its flow rate is held at 70 ml/min by a Brooks flow control valve.

Just above the fritted disc is the sample inlet. It is connected

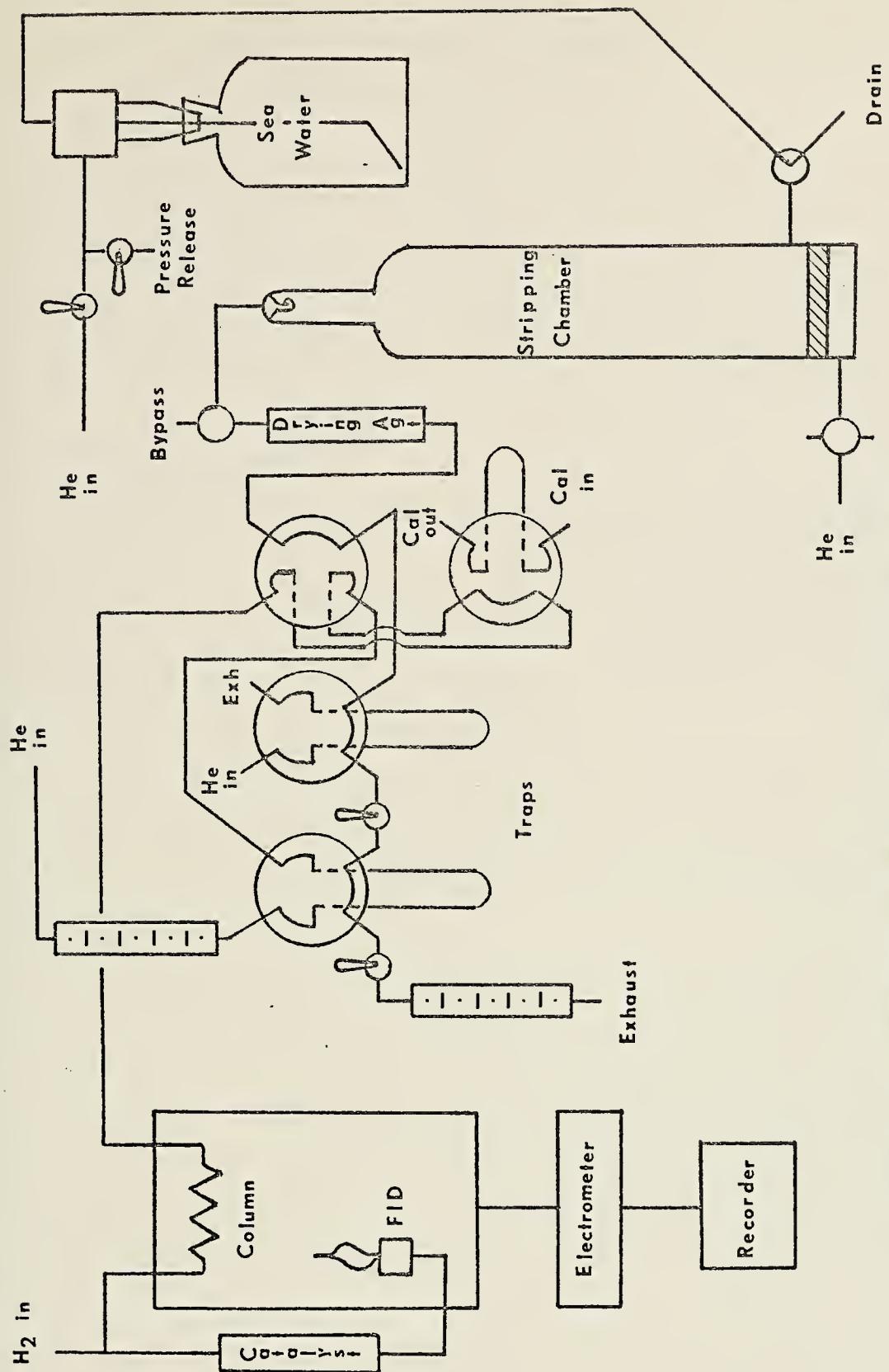


Figure 1. Schematic diagram of the dissolved gas analysis system. System is shown in the analyze position.

to one port of a Y-type stopcock. This stopcock allows the operator to purge the line prior to sample transfer, transfer the sample, and later, drain the chamber.

The neck of the chamber is fitted with a ground glass joint. This provides an opening for cleaning and a place to insert the magnetic stirring bar. At the top of the neck is a modified Kjeldahl tip. This breaks any bubbles and thus keeps water from being carried into the rest of the system.

3. Trapping Columns

The traps were constructed of 3/16" stainless steel tubing. The first trap was packed with ten inches of 60/80 mesh activated alumina and the second with a mixture containing $\frac{1}{4}$ activated charcoal and 3/4 30/60 mesh 5A molecular sieve. They were connected across the sample loop ports of Perkin-Elmer gas sampling valves (See Figure 2).

Toggle valves were placed on either side of the activated charcoal/ molecular sieve trap. These allowed the trap to be isolated while holding a sample and thus precluded any gas leakage prior to analysis.

4. Backflush Lines

A constant flow of helium had to be maintained across the column in the chromatograph. This was accomplished in that in the trap position, the carrier gas was passing through the gas sampling valve for the activated charcoal/molecular sieve trap to the chromatograph. In the analyze position, the carrier gas backflushed the trapped sample gases to the chromatograph. The flow rate was maintained at 30 ml/min by a Brooks flow control valve.

After the analysis was complete, the activated alumina trap was backflushed with line helium to remove trapped gases. Since this trap was

not analyzed, a cap was placed on the exhaust port when the valve was in the trap position, in order to conserve helium and maintain pressure for stripping. It also kept water from backing up into the system.

B. CALIBRATION SYSTEM

In order to have consistent calibrations, a gas sampling valve with a 1 ml. sample loop was incorporated into the system. This enabled a known gas mixture to be introduced into the trapping system or directly into the chromatograph. A diagram showing the calibration system is shown in Figure 2.

C. GAS CHROMATOGRAPH

A Varian Aerograph Model 600C gas chromatograph was obtained on loan from the Department of Physics and Chemistry. It was equipped with a flame ionization detector (FID), a tube-type electrometer, and a Model 328 isothermal temperature controller.

1. Modifications

The original column was replaced with a four foot, 3/16", stainless steel column packed with 30/60 mesh 5A molecular sieve. Upon conditioning (heating for two hours at 150 °C in a helium atmosphere), this column provided adequate separation of carbon monoxide and methane.

The normal carrier-gas port and injection port on the gas chromatograph were bypassed since the already mixed unknown and carrier gas were entering from the traps, i.e., there was a direct connection of the line from the activated charcoal/molecular sieve trap to the column.

The hydrogen port was also sealed since the hydrogen was introduced prior to a catalyst tube.

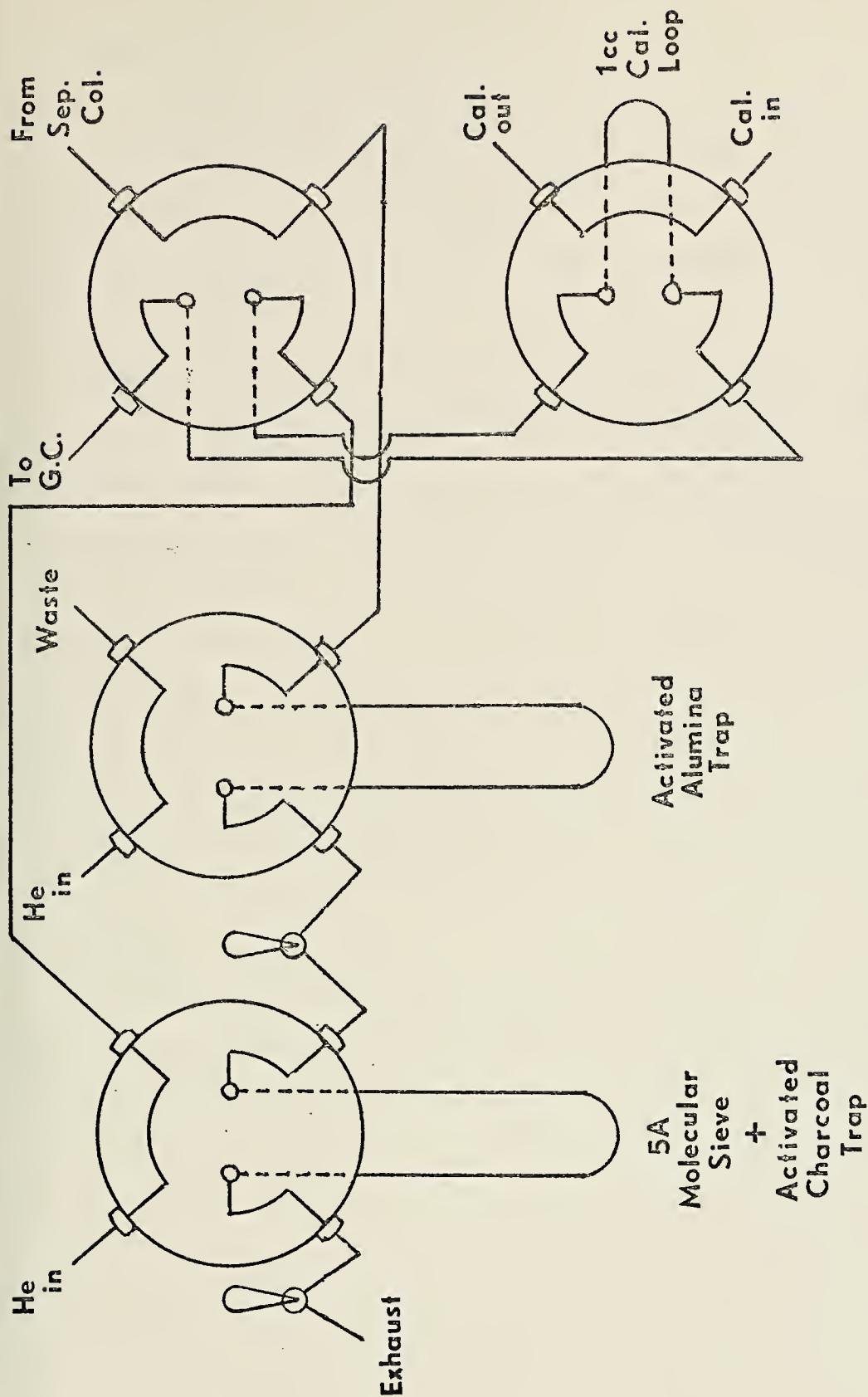


Figure 2. Detailed schematic diagram of the calibration system and traps. System is shown in the trap position with a calibration sample being injected into the chromatograph.

D. CATALYST FURNACE

The flame ionization detector is not sensitive to carbon monoxide. Therefore, the carbon monoxide must be converted to methane prior to analysis. This is accomplished by passing the carbon monoxide over a heated nickel catalyst in the presence of hydrogen (Porter and Vollman, 1962).

The catalyst furnace was placed between the analytical column and the detector. It consisted of a 10 cm., $\frac{1}{4}$ ", stainless steel tube in which the catalyst was placed. An aluminum block 7.5 cm. long and 2.5 cm. in diameter was placed around this tube as a heat sink. This assembly was placed flush inside a threaded quartz tube which was wound with a Nichrome wire heater. To prevent fire and provide additional insulation, the heater was placed in a Transite box which was mounted on the side of the chromatograph. The temperature was monitored with a five ohm iron-constantan thermocouple attached to the auxillary terminal of the temperature controller. Temperature control was regulated with a Variac controller.

E. RECORDER

A Varian Model G-14 strip chart recorder was used to record the output of the electrometer. It was set for 1 mv. full scale deflection and run at a chart speed of 1 in./min.

III. EXPERIMENTAL METHODS

A. SAMPLE COLLECTION

Five nearshore stations were taken in kelp beds around the Monterey Peninsula as described in Table 1 and Figure 3. The stations varied from a relatively calm area at Del Monte Beach to areas of extreme turbulence on the exposed coast at Point Pinos and Point Joe. These stations were occupied periodically from 31 October 1972 to 15 December 1972.

Sampling in the nearshore areas was done from a forty foot boat. In order to minimize contamination from the exhaust, the boat was allowed to drift to a stop in the kelp beds prior to sampling. Samples were taken in a Van Dorn type bottle that was rigged with hand lines for lowering and tripping.

Samples were collected in 500 ml. reagent bottles by filling from the bottom with a surgical rubber tube. The bottles were then sealed with a standard taper ground glass stopper. This reduced the possibility of air contamination and outgassing from the sample. If the samples were to be stored before analysis, a small amount of sodium azide (NaN_3) was added to kill any organisms that might alter the dissolved gas concentrations.

In order to have open ocean values to compare with those found in the nearshore habitats, a deep station was taken in the Monterey Canyon aboard R/V Acania (see Table 1). Samples were taken at depths of 0, 5, 15, 30, 50, 75, 100, 200, 500, 700, and 1000 meters using standard Nansen bottles. In addition, a cast was made at 0, 5, 12, 18, and 45 meters for determination of primary productivity (see Rowney, 1973).

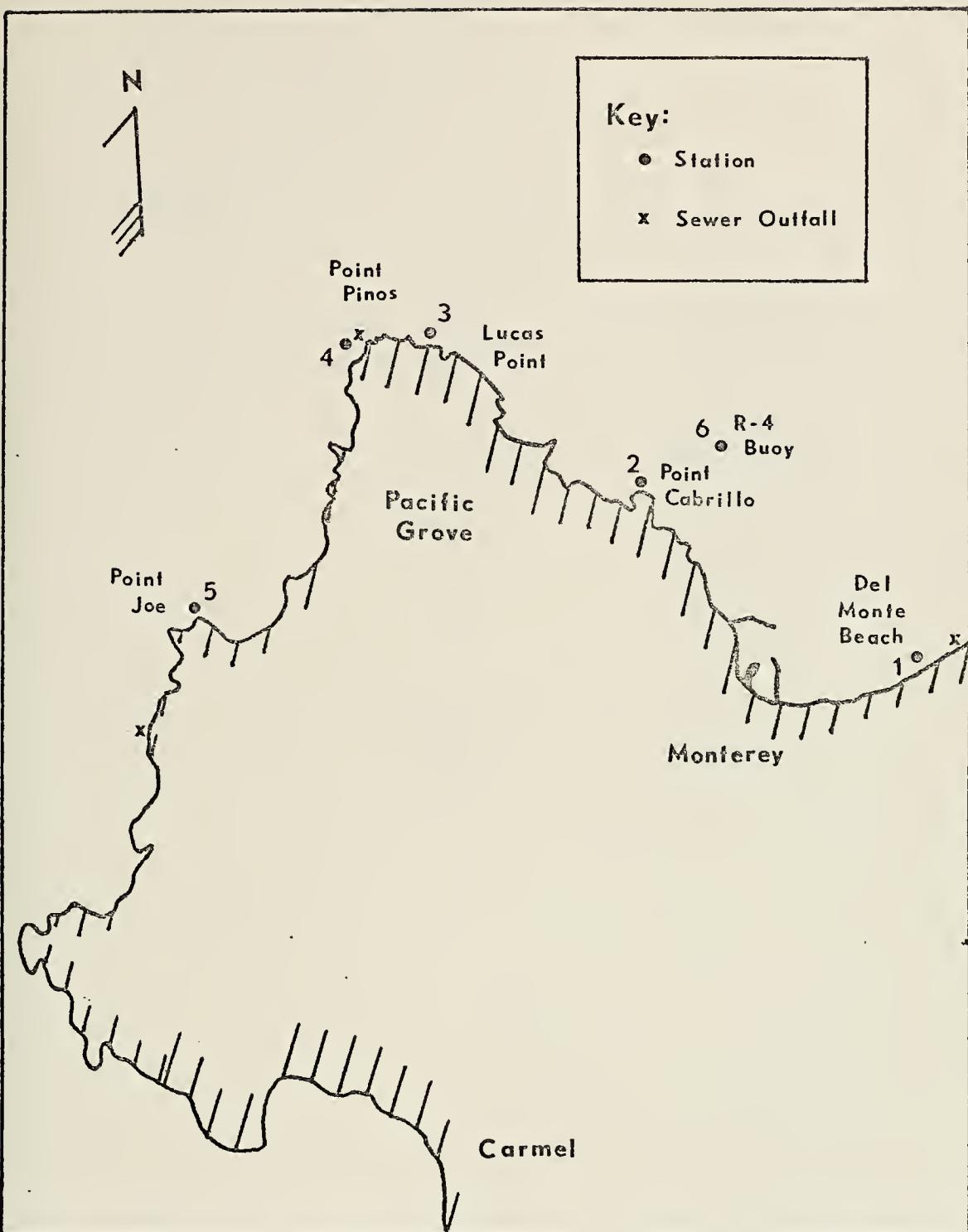


Figure 3. Chart showing location of nearshore stations.

Table 1

Seawater Sampling Stations

<u>Station Number</u>	<u>Name</u>	<u>Positions</u>
1	Del Monte Beach	$36^{\circ} 36.3'$ N $121^{\circ} 52.4'$ W
2	Point Cabrillo	$36^{\circ} 37.4'$ N $121^{\circ} 54.1'$ W
3	Point Pinos North	$36^{\circ} 38.3'$ N $121^{\circ} 55.4'$ W
4	Point Pinos South	$36^{\circ} 38.3'$ N $121^{\circ} 56.3'$ W
5	Point Joe	$36^{\circ} 36.8'$ N $121^{\circ} 57.3'$ W
6	R-4 Bell Buoy	$36^{\circ} 37.5'$ N $121^{\circ} 54.3'$ W
7	Monterey Canyon	$36^{\circ} 44.3'$ N $122^{\circ} 07.2'$ W

Transects were taken across Del Monte Beach, from Del Monte Beach to the R-4 buoy, and from Point Cabrillo to the R-4 buoy. These were taken to determine any gradients present.

B. SAMPLE ANALYSIS¹

The sample bottle was connected to the system by a ground glass joint. Helium was introduced above the sample forcing the water through a stainless steel tube to the drain valve. The first few milliliters were discarded to ensure flushing of the transfer line. The stripping chamber, which had been previously drained to the 500 ml. mark, was then filled to the zero mark. This method of transfer minimized contact with the atmosphere and thus reduced the chance of contamination.

"Purge helium" was introduced below the fritted disc and allowed to bubble through the sample for twelve minutes. A magnetic stirring bar in the chamber increased the residence time of the bubbles in the sample. After twelve minutes, no measurable amounts of carbon monoxide or methane were left in the sample and as prior tests showed, no leakage from the cold traps had occurred. Before reaching the traps, the gas passed through a magnesium perchlorate drying agent to remove any water vapor.

Trapping of the gases took place on two series-connected cold traps. They were both maintained at -77°C with an acetone-dry ice bath, in a Dewar flask. The first trap was an activated alumina column which removed any hydrocarbons higher than methane as well as carbon dioxide. The second, a mixture of 1/4 activated charcoal and 3/4 molecular sieve, removed carbon monoxide, methane, and air. In the twelve minute purging time, most

¹ A complete, step-by-step, analysis check list is included in Appendix A.

of the air was bled off the trap and only a small amount of residual oxygen remained. This proved to be no problem since the analytical column in the gas chromatograph provided sufficient separation of these gases.

After the twelve minutes of purging time, isolation valves on either side of the activated charcoal/molecular sieve trap were closed. The acetone-dry ice bath was removed and replaced with one of boiling water. The trap was now backflushed into the analytical column in the gas chromatograph. The activated alumina trap was backflushed to the atmosphere since its contents were not of importance to this experiment.

Retention times for a column temperature of 48°C and a flow rate of 30 ml/min were: air, 1.5 min.; methane, 3.0 min.; carbon monoxide, 4.25 min. Analysis was complete seven minutes after injection of the sample into the gas chromatograph.

To ensure complete conversion of carbon monoxide to methane, the catalyst furnace was held between 300-320 °C. This also reduced tailing on the carbon monoxide peak.

Figure 4 shows a chromatogram from a sample analysis. The concentration of each gas is proportional to the area under its respective peak. In order for these areas to be meaningful, a method of calibration was needed.

To accomplish this, a calibration gas is used. A calibrated gas mixture of 70.3 ppm methane and 71.1 ppm carbon monoxide in air was obtained from the Naval Research Laboratory, Washington D.C. Preliminary work showed no measurable difference in the results when the gas was introduced into the traps or directly into the gas chromatograph, therefore direct injection was used to conserve time.

Figure 5 shows the results of a calibration run. Now, the response of the system to a known concentration of gas was known and the unknown

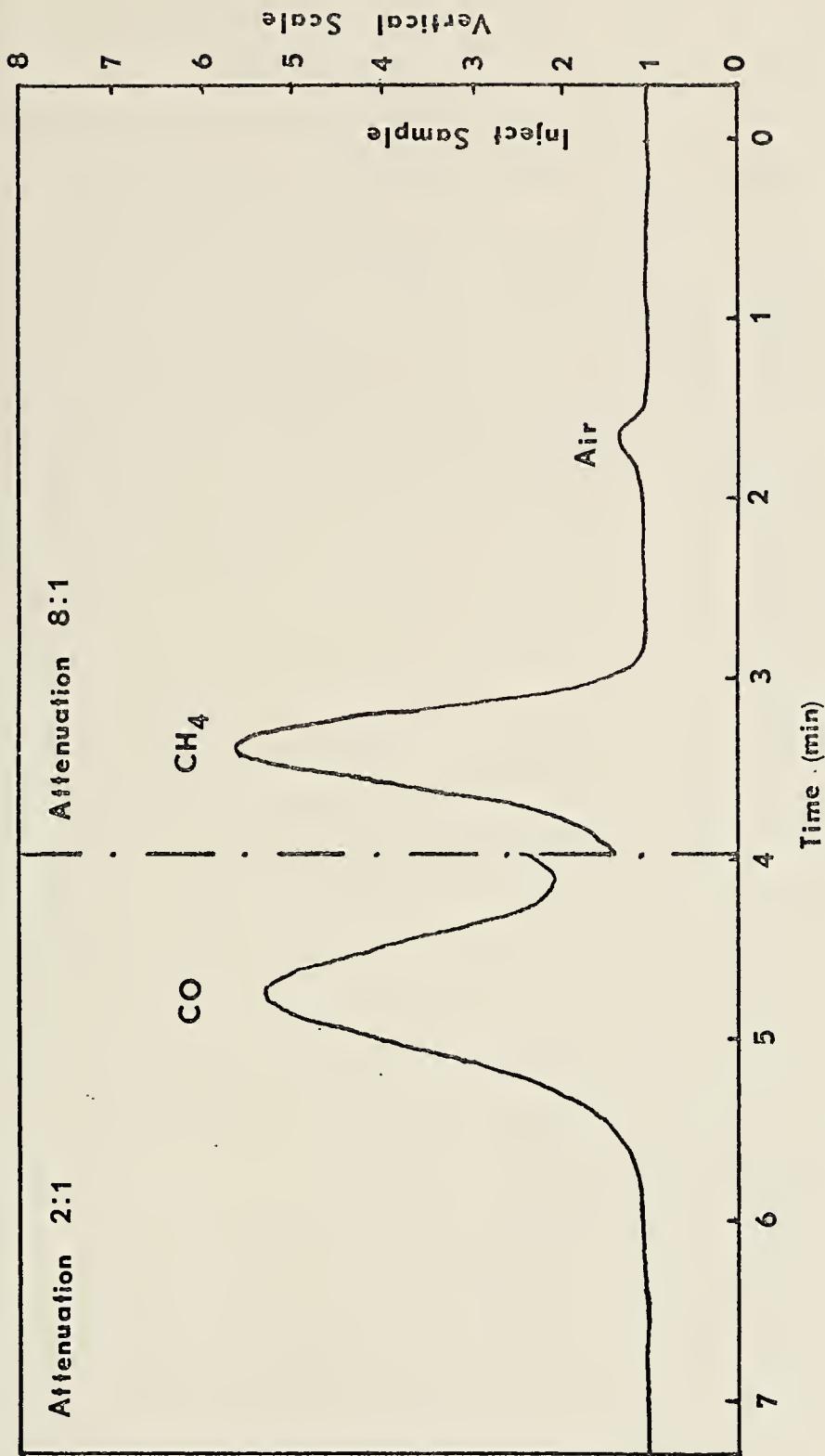


Figure 4. Chromatogram from a sample run. Signals from the electrometer have been attenuated 2:1 for CO and 8:1 for CH₄.

concentrations could be determined by comparing areas. Calibrations were run periodically along with the samples. In addition to calibrating the system, these calibration checks were also used to determine if the catalytic conversion of carbon monoxide to methane was complete.

After the sample was run, the area under each peak was determined in the following manner. The height of the peak and the width at half height were measured with a set of dial calipers. These values were multiplied together to get the raw area; a range and attenuation factor from the electrometer setting was applied to get the effective area. These calculations were carried out on the IBM 360/67 computer. A copy of the program is included in Appendix C.

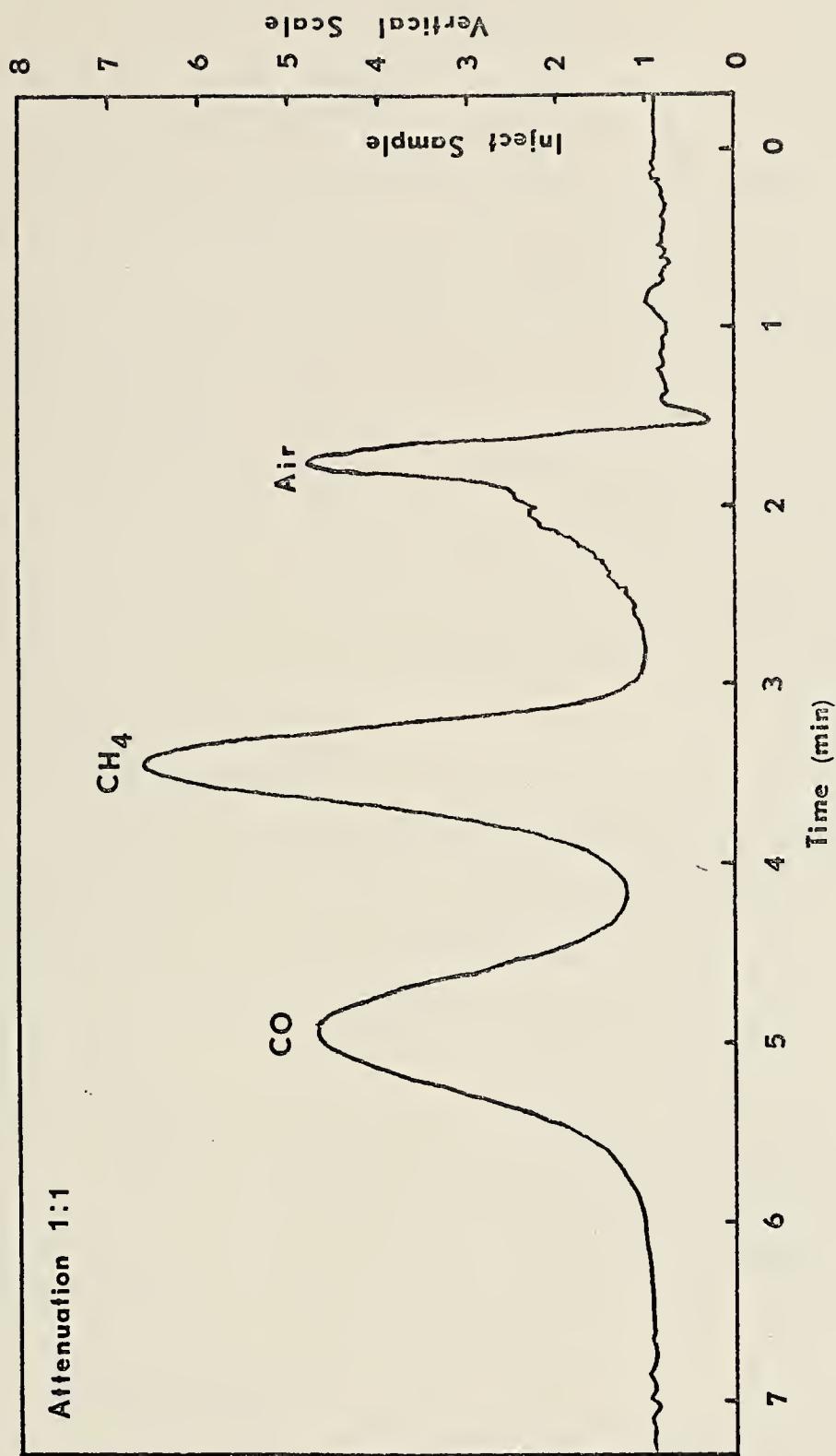


Figure 5. Chromatogram from a calibration run.

IV. RESULTS

A. OPEN OCEAN DEEP STATION

Station seven (Table 1) was located on the axis of the Monterey Submarine Canyon in 1375 meters of water. An eleven bottle cast to 1000 meters was taken. The profiles from this cast are shown in Figures 6 and 7 and tabulated in Tables 2 and 3.

1. Methane

The methane profile (Figure 6 and Table 2) shows a surface concentration of 1.1×10^{-4} ml/l. This value increases with depth to 50 meters. From 50 to 100 meters it decreases slightly and then more rapidly to 200 meters. It then decreases linearly with depth at a rate of 0.125×10^{-4} ml/l/100 m to 1000 meters.

2. Carbon Monoxide

Figure 7 and Table 3 shows a carbon monoxide concentration of 0.81×10^{-4} ml/l at the surface. This value increases sharply to almost 2×10^{-4} ml/l at 15 meters and then decreases to near surface values at 100 meters. From 100 to 1000 meters, the concentration decreases to only trace amounts.

3. Primary Productivity

Primary productivity measurements made by Rowney (1973) at the same station are shown in Figure 8. Note the relatively high values in the upper 12 meters.

B. TEMPORAL STUDIES

A temporal study of carbon monoxide and methane was conducted in each of the five nearshore habitats. The plots of concentration versus time

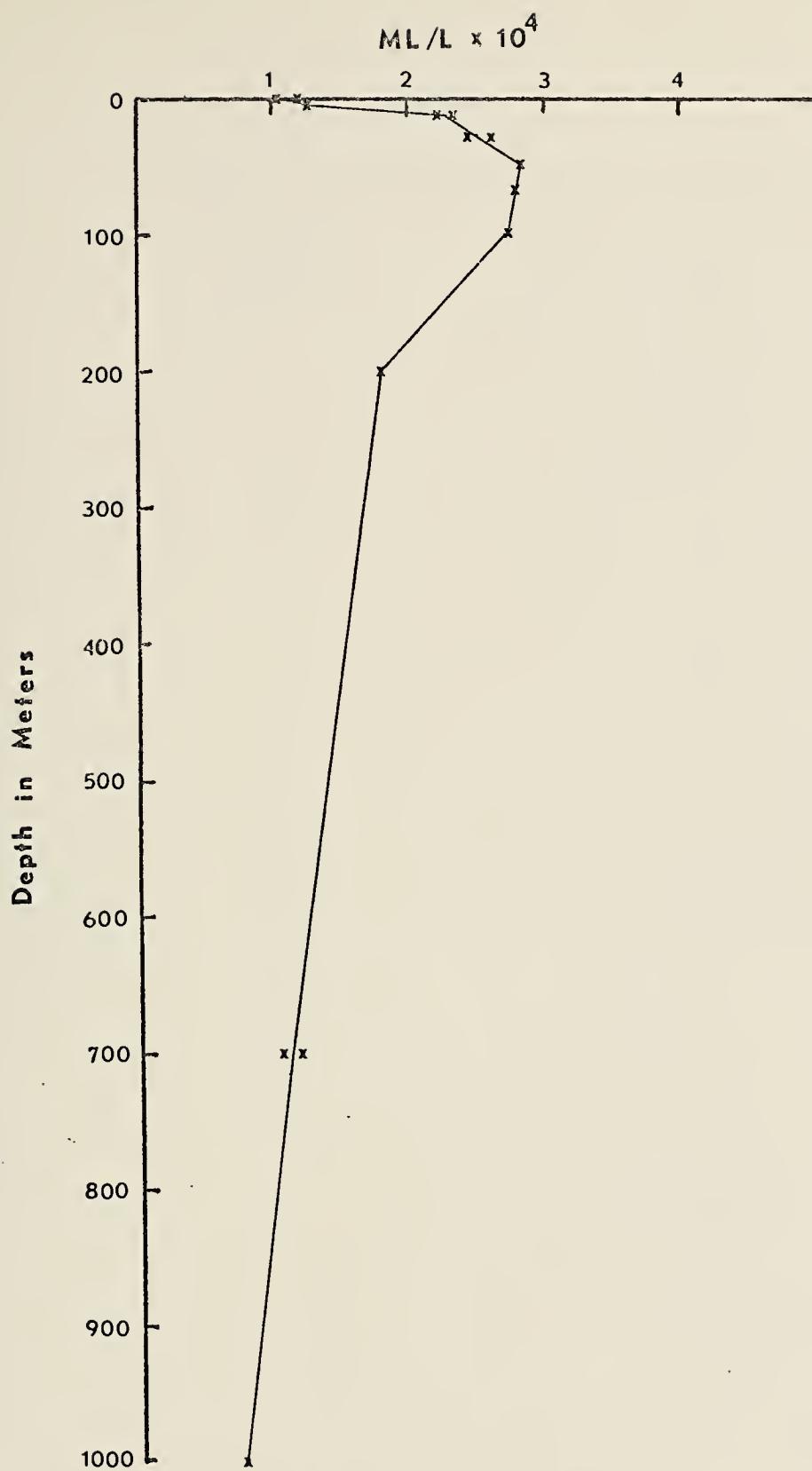


Figure 6. Vertical distribution of CH_4 at the Deep Ocean Station in Monterey Canyon.

Table 2

Methane Concentrations in Monterey Canyon

<u>Depth (meters)</u>	Methane Concentration (ml/l)
0	1.013×10^{-4}
0	1.184×10^{-4}
5	1.252×10^{-4}
15	2.342×10^{-4}
15	2.205×10^{-4}
30	2.433×10^{-4}
30	2.632×10^{-4}
50	2.847×10^{-4}
75	2.783×10^{-4}
100	2.752×10^{-4}
200	1.797×10^{-4}
200	1.759×10^{-4}
700	1.037×10^{-4}
1000	0.751×10^{-4}

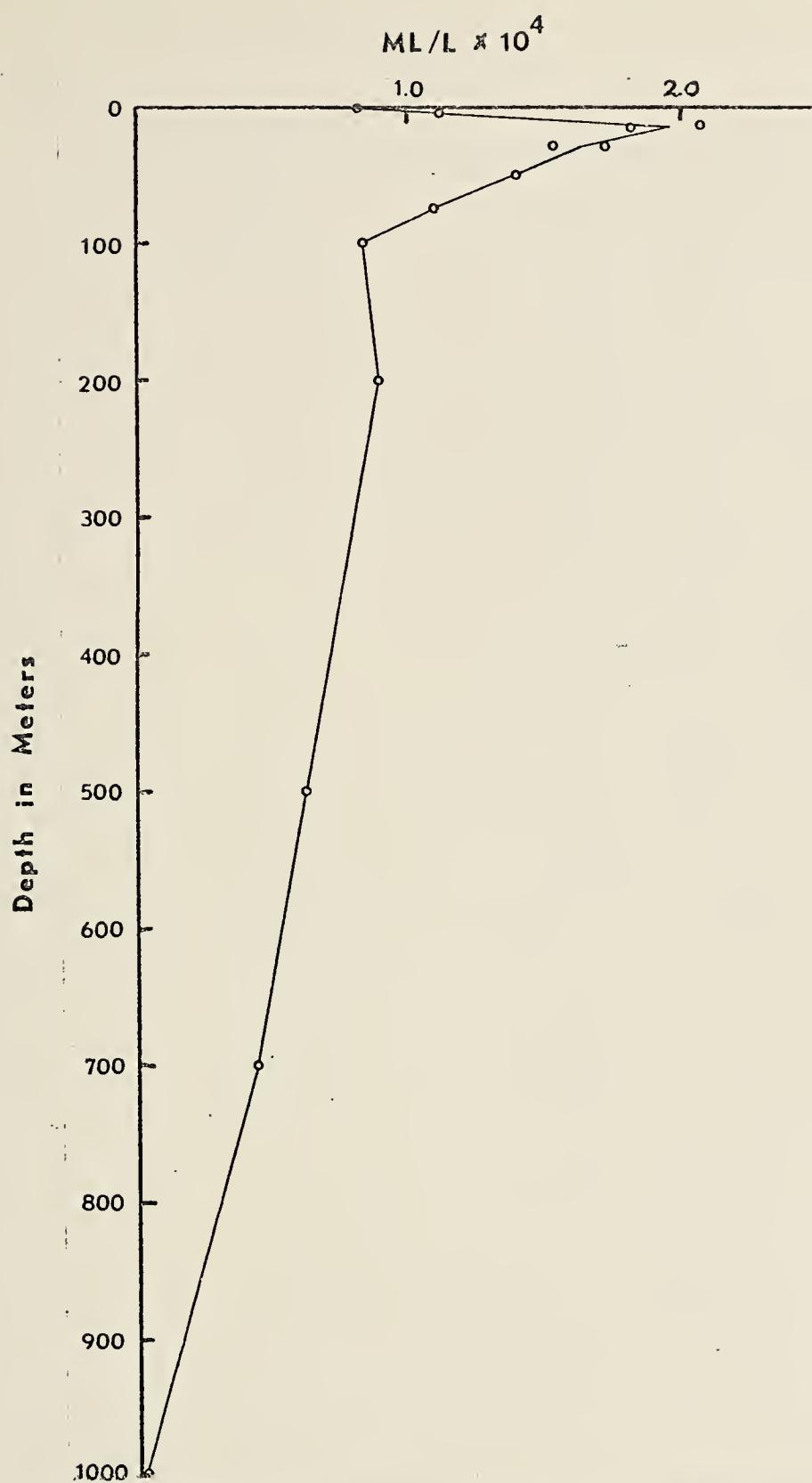


Figure 7. Vertical distribution of CO₂ at the Deep Ocean Station in Monterey Canyon.

Table 3

Carbon Monoxide Concentrations in Monterey Canyon

<u>Depth (meters)</u>	Carbon Monoxide Concentration (ml/l)
0	0.793×10^{-4}
0	0.836×10^{-4}
5	1.230×10^{-4}
15	2.144×10^{-4}
15	1.818×10^{-4}
30	1.522×10^{-4}
30	1.729×10^{-4}
50	1.391×10^{-4}
75	1.086×10^{-4}
100	0.759×10^{-4}
200	0.836×10^{-4}
200	0.859×10^{-4}
500	0.619×10^{-4}
700	0.436×10^{-4}
1000	trace

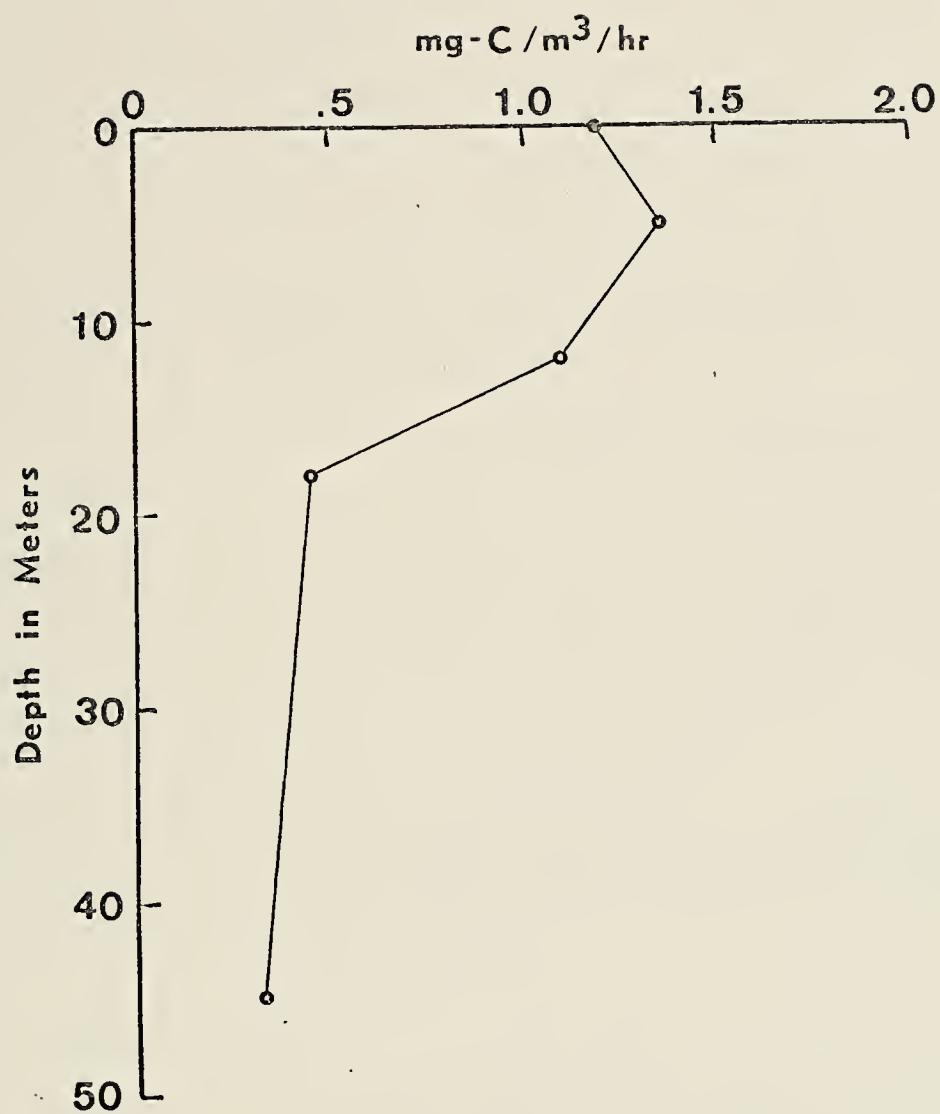


Figure 8. Vertical distribution of Primary Productivity in the upper 50 meters of the Deep Ocean Station in Monterey Canyon (from Rowney, 1973).

are found in Figures 9 through 13 and are tabulated in Tables 4 through 8.

Environmental conditions were as follows. The period prior to 31 October 1972 had been clear for several weeks. During the periods 4-17 November and 3-8 December 1972, overcast skies with heavy rains were predominant. The rest of the time, clear skies prevailed. Heavy swell accompanied the periods of rain. This made it impossible to sample stations 3, 4, and 5 during these periods, since it was too rough for the forty foot boat. The surface temperature (Figure 14) decreased from 15⁰C at the start of the study to 10⁰C at its conclusion.

All five stations showed high methane concentrations on the first day. These values dropped sharply after the onset of the first storm. Following this storm, the concentrations began to rise, but were again decreased after the second storm. Del Monte Beach showed another increase after this second storm that was not noted at the other stations.

Carbon monoxide did not seem to be affected by the rains as much as methane. It was uniformly low at the start of the study with a general increase throughout the time period. Again, Del Monte Beach showed more variability than the other stations.

C. GRADIENT ANALYSIS

The average of all measurements at each nearshore station was computed and plotted (Figure 15). This provided a gradient analysis for comparison with primary productivity, chlorophyll and nutrient analyses (Rowney 1973).

The methane gradient decreased from Del Monte Beach to Point Cabrillo. It then reversed and showed an increase to the highest value at Point Joe.

Carbon monoxide showed a different pattern. It alternated between decreasing and increasing gradients with a slight overall trend towards higher concentrations at Point Joe.

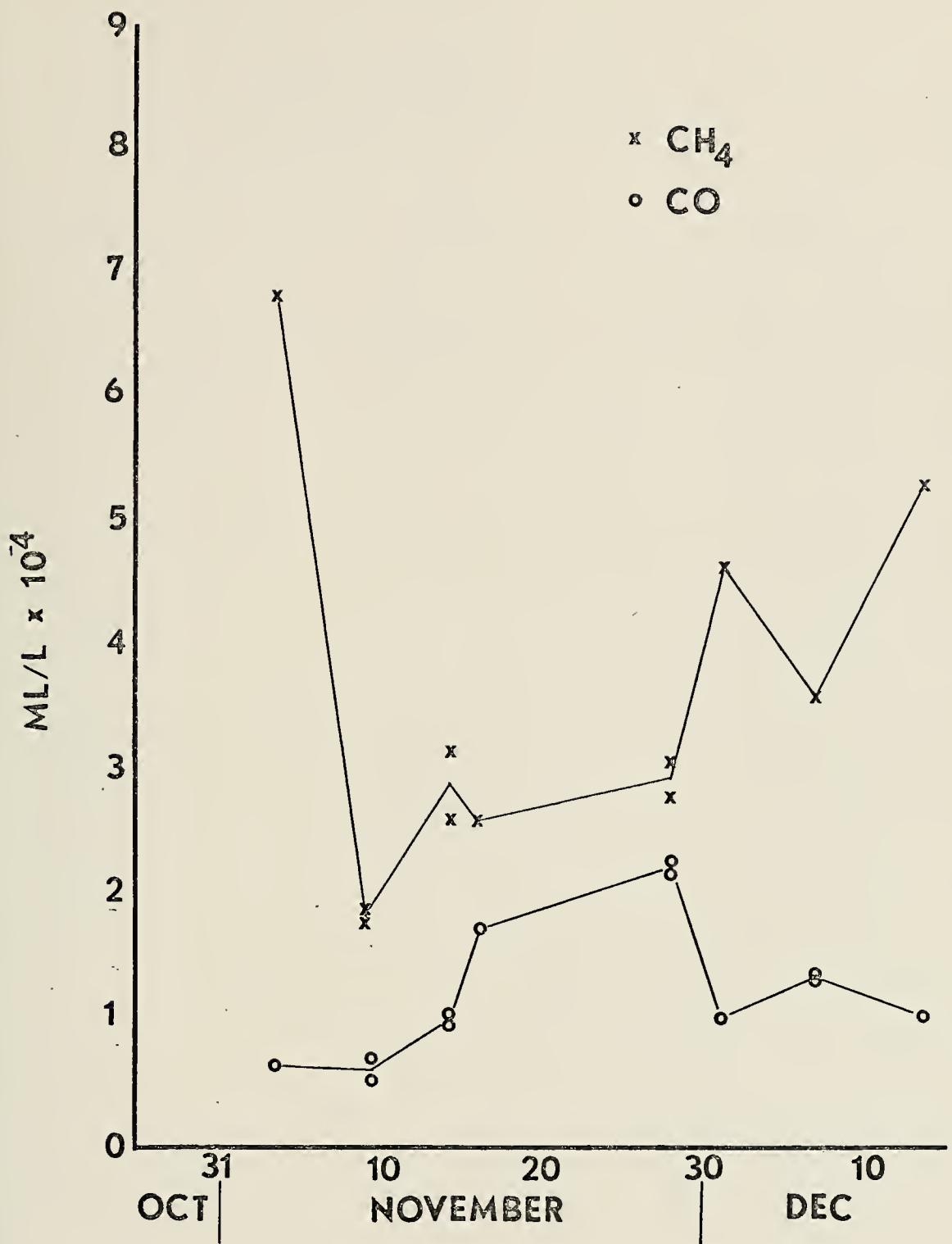


Figure 9. CH_4 and CO concentrations in the surface waters at Del Monte Beach.

Table 4
 Methane and Carbon Monoxide
 Concentrations at Del Monte Beach

<u>Date</u>	<u>Methane</u> (ml/l)	<u>Carbon Monoxide</u> (ml/l)
11-03-72	6.829×10^{-4}	0.663×10^{-4}
11-09-72	1.842×10^{-4}	0.546×10^{-4}
11-09-72	1.959×10^{-4}	0.735×10^{-4}
11-14-72	2.648×10^{-4}	0.990×10^{-4}
11-14-72	3.187×10^{-4}	1.088×10^{-4}
11-16-72	2.631×10^{-4}	1.763×10^{-4}
11-28-72	3.102×10^{-4}	2.208×10^{-4}
11-28-72	2.837×10^{-4}	2.318×10^{-4}
12-01-72	4.682×10^{-4}	1.039×10^{-4}
12-07-72	3.669×10^{-4}	1.408×10^{-4}
12-07-72	3.547×10^{-4}	1.345×10^{-4}
12-14-72	5.358×10^{-4}	1.046×10^{-4}
12-14-72	5.277×10^{-4}	1.082×10^{-4}

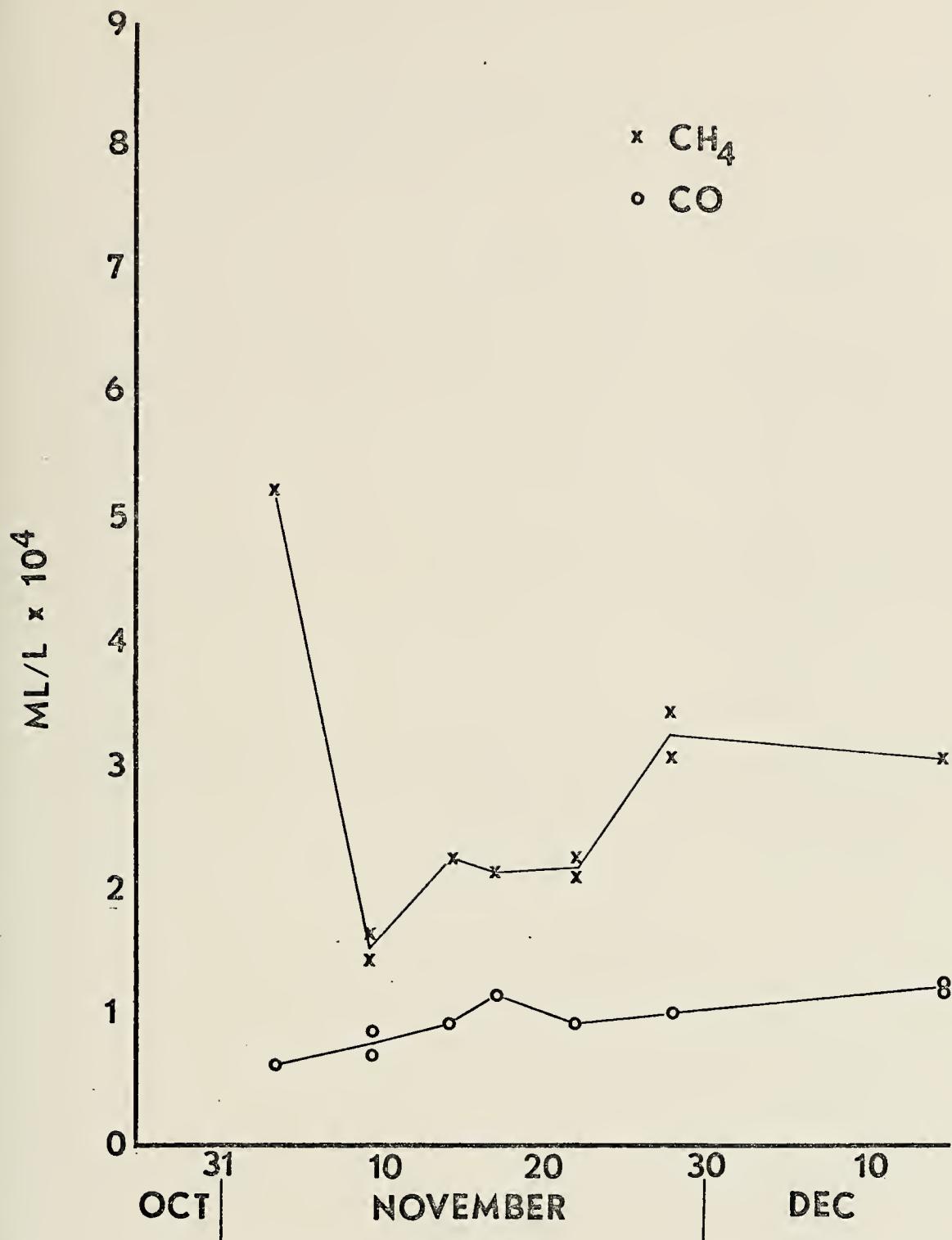


Figure 10. CH_4 and CO concentrations in the surface waters at Point Cabrillo.

Table 5
 Methane and Carbon Monoxide
 Concentrations at Point Cabrillo

Date	<u>Methane</u> (ml/l)	<u>Carbon Monoxide</u> (ml/l)
11-03-72	5.273×10^{-4}	0.677×10^{-4}
11-09-72	1.714×10^{-4}	0.942×10^{-4}
11-09-72	1.499×10^{-4}	0.733×10^{-4}
11-14-72	2.323×10^{-4}	0.983×10^{-4}
11-17-72	2.219×10^{-4}	1.209×10^{-4}
11-21-72	2.323×10^{-4}	0.955×10^{-4}
11-21-72	2.152×10^{-4}	1.006×10^{-4}
11-28-72	3.106×10^{-4}	1.066×10^{-4}
11-28-72	3.478×10^{-4}	1.097×10^{-4}
12-15-72	3.117×10^{-4}	1.242×10^{-4}
12-15-72	3.096×10^{-4}	1.305×10^{-4}

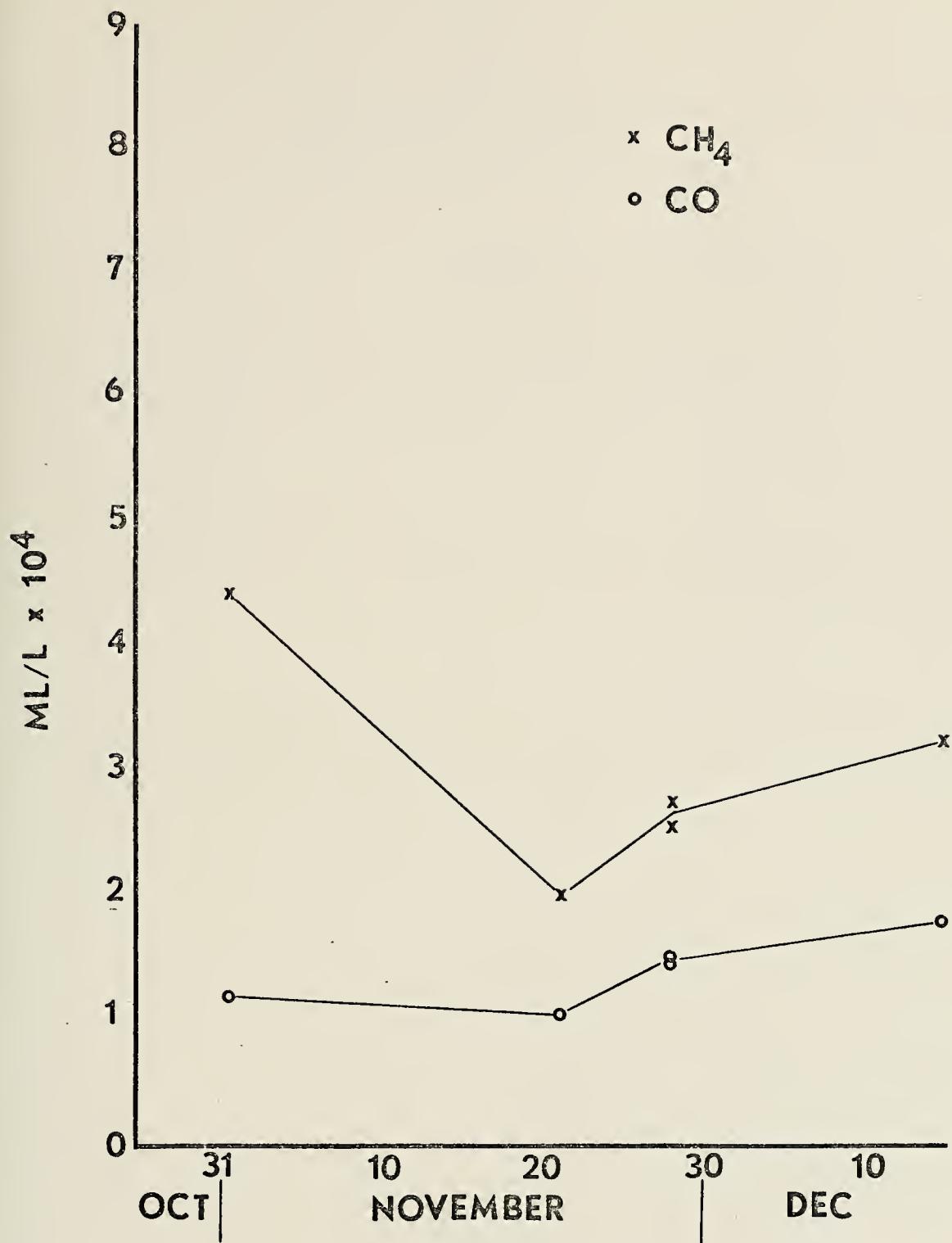


Figure 11. CH_4 and CO concentrations in the surface waters at Point Pinos North.

Table 6

Methane and Carbon Monoxide
 Concentrations at Point Pinos North

<u>Date</u>	<u>Methane</u> <u>(ml/1)</u>	<u>Carbon Monoxide</u> <u>(ml/1)</u>
10-31-72	4.437×10^{-4}	1.203×10^{-4}
11-21-72	2.014×10^{-4}	1.085×10^{-4}
11-28-72	2.791×10^{-4}	1.531×10^{-4}
11-28-72	2.587×10^{-4}	1.467×10^{-4}
12-15-72	3.267×10^{-4}	1.833×10^{-4}

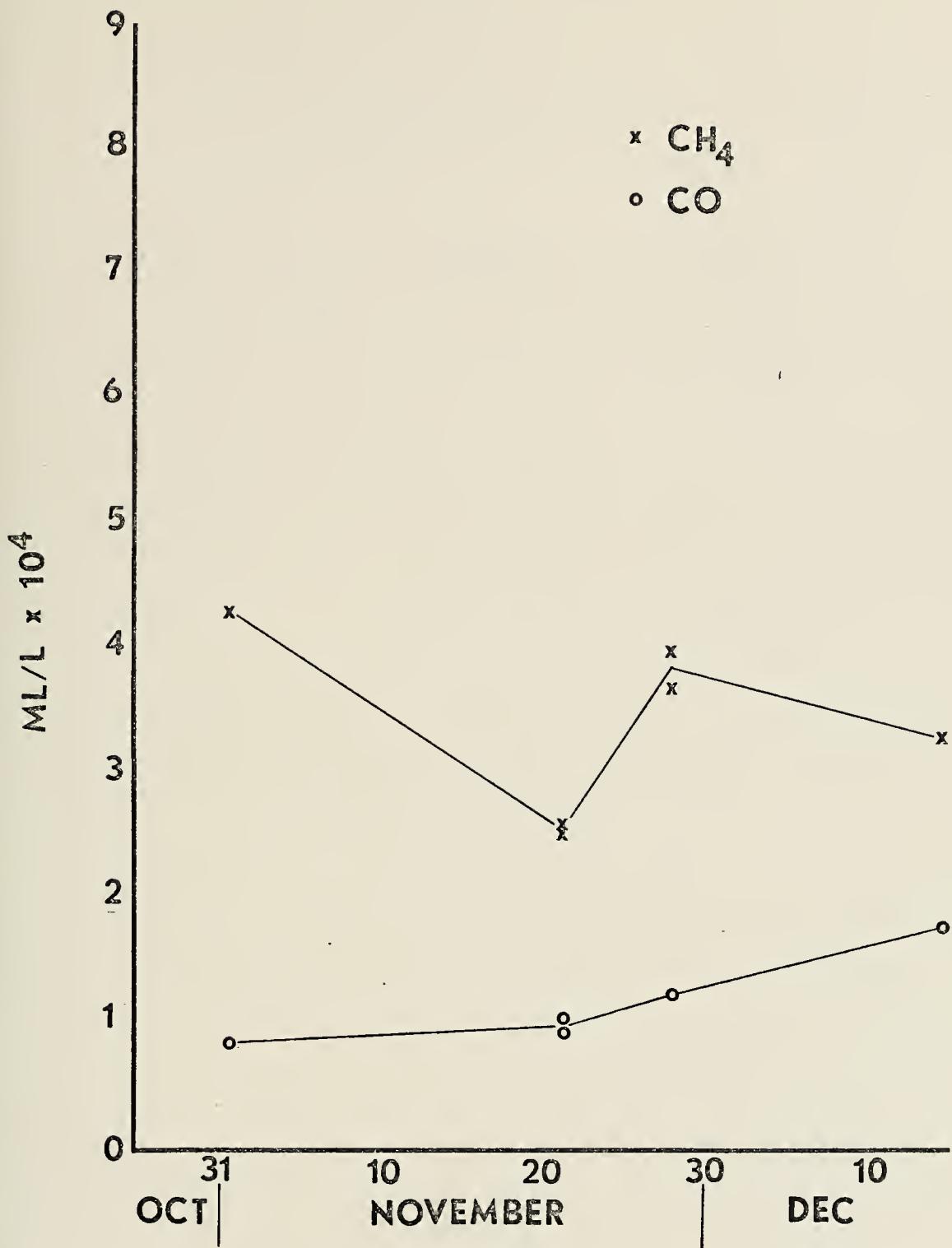


Figure 12. CH_4 and CO concentrations in the surface waters at Point Pinos South.

Table 7
 Methane and Carbon Monoxide
 Concentrations at Point Pinos South

<u>Date</u>	<u>Methane</u> (ml/l)	<u>Carbon Monoxide</u> (ml/l)
10-31-72	4.317×10^{-4}	0.875×10^{-4}
11-21-72	2.527×10^{-4}	0.952×10^{-4}
11-21-72	2.635×10^{-4}	1.061×10^{-4}
11-28-72	3.998×10^{-4}	1.288×10^{-4}
11-28-72	3.694×10^{-4}	1.251×10^{-4}
12-15-72	3.481×10^{-4}	1.798×10^{-4}

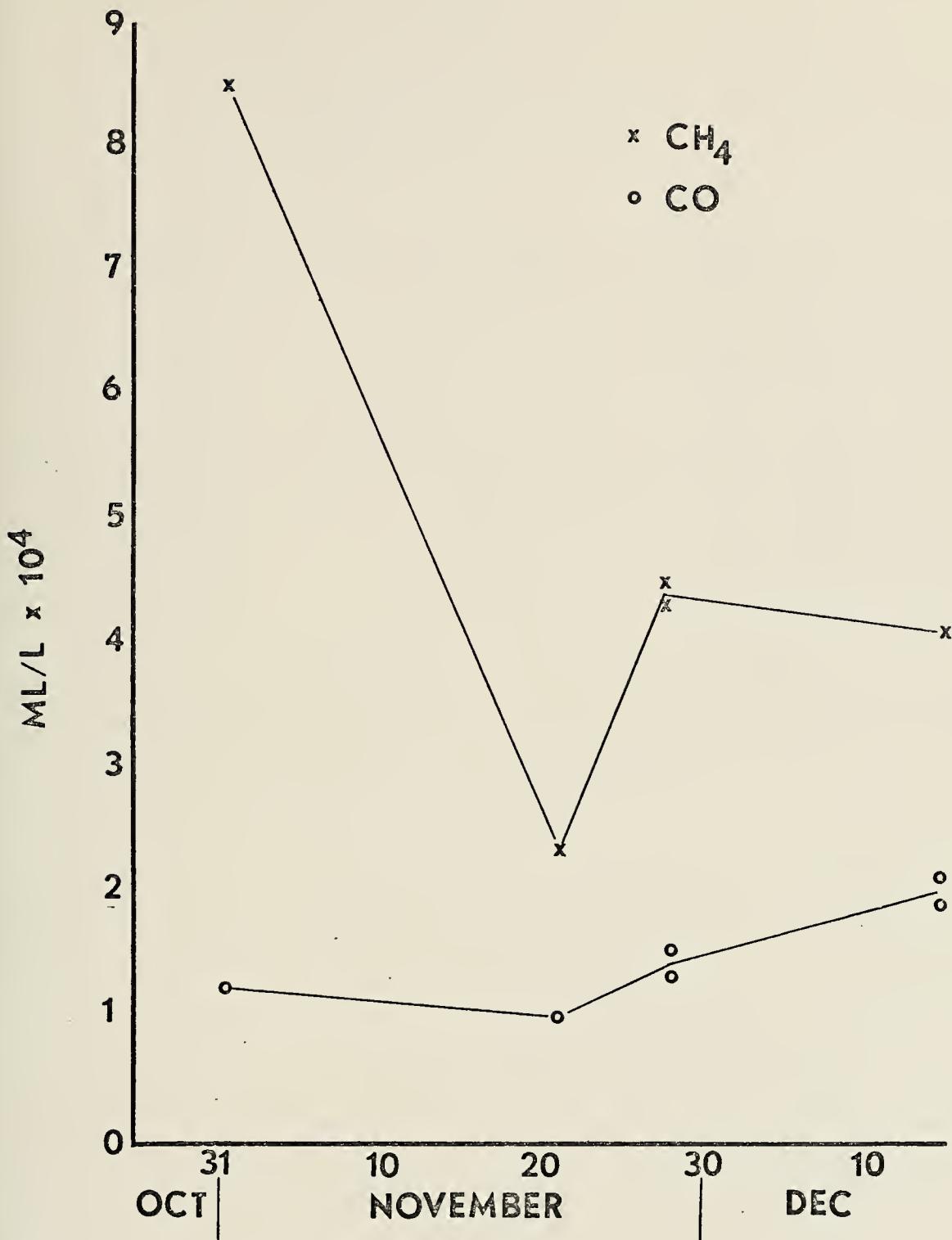


Figure 13. CH_4 and CO concentrations in the surface waters at Point Joe.

Table 8
 Methane and Carbon Monoxide
 Concentrations at Point Joe

<u>Date</u>	<u>Methane</u> <u>(ml/l)</u>	<u>Carbon Monoxide</u> <u>(ml/l)</u>
10-31-72	8.524×10^{-4}	1.275×10^{-4}
11-21-72	2.300×10^{-4}	1.034×10^{-4}
11-28-72	4.267×10^{-4}	1.548×10^{-4}
11-28-72	4.501×10^{-4}	1.344×10^{-4}
12-15-72	4.088×10^{-4}	2.109×10^{-4}
12-15-72	4.073×10^{-4}	1.903×10^{-4}

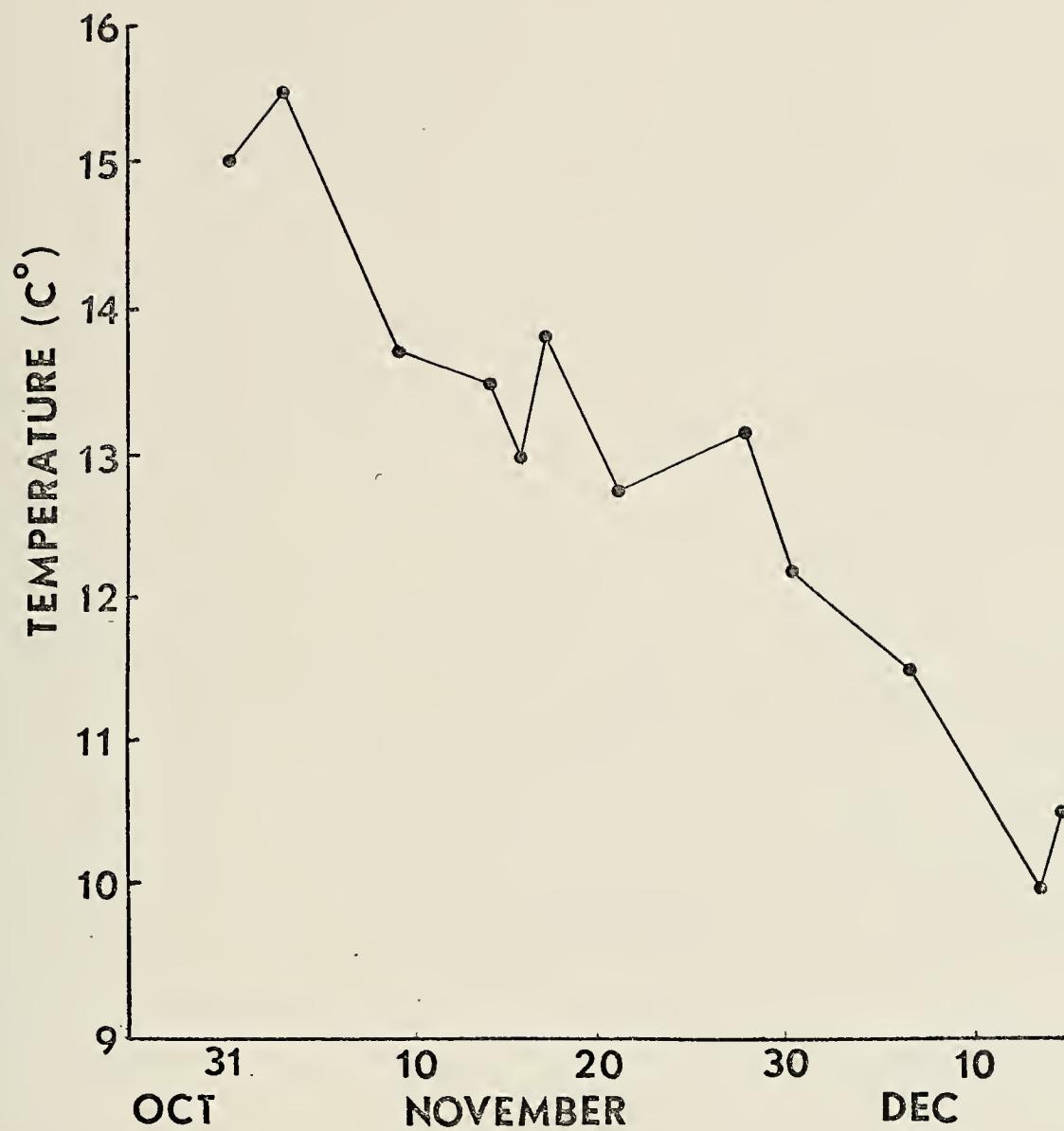


Figure 14. Surface temperature.

D. TRANSECTS

The transect across Del Monte Beach proved most interesting. It started 400 yards east of the Monterey sewage disposal plant outfall and proceeded towards the harbor. Figure 16 and Table 9 show the results of this transect. The initial values of methane and carbon monoxide were low and then increased sharply at the "boil" above the outfall. Concentrations dropped back down west of the "boil" and then started to increase again as the stations got closer to Monterey harbor.

The transects from Del Monte Beach to the R-4 buoy (Figure 17 and Table 10) showed a wide variation in concentrations. On 1 December 1972, the methane decreased slightly to the first station beyond the kelp bed and then increased out to the buoy. Carbon monoxide increased slightly beyond the kelp and then remained fairly constant. On 14 December 1972, the methane followed the same pattern, but the drop in concentration from the kelp bed to the first open water station was much greater. Carbon monoxide seemed to decrease slightly beyond the kelp bed and then build up seaward.

The transect from Point Cabrillo to the R-4 buoy (Figure 18 and Table 11) showed very little change in either gas. Methane increased by $.1 \times 10^{-4}$ ml/l from 3.1×10^{-4} ml/l and carbon monoxide decreased $.4 \times 10^{-4}$ ml/l from 1.3×10^{-4} ml/l.

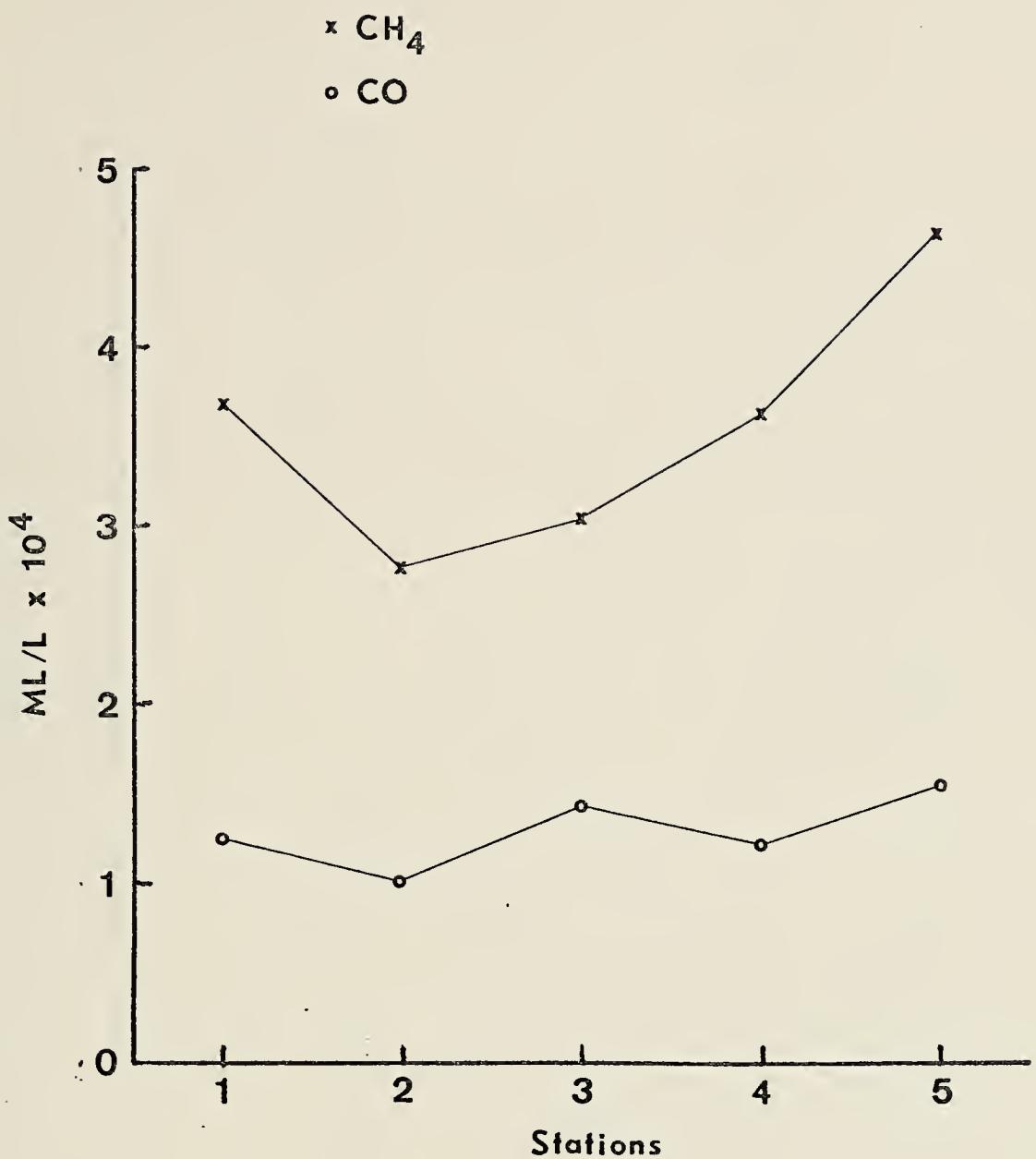


Figure 15. CH_4 and CO gradients between the five nearshore stations.

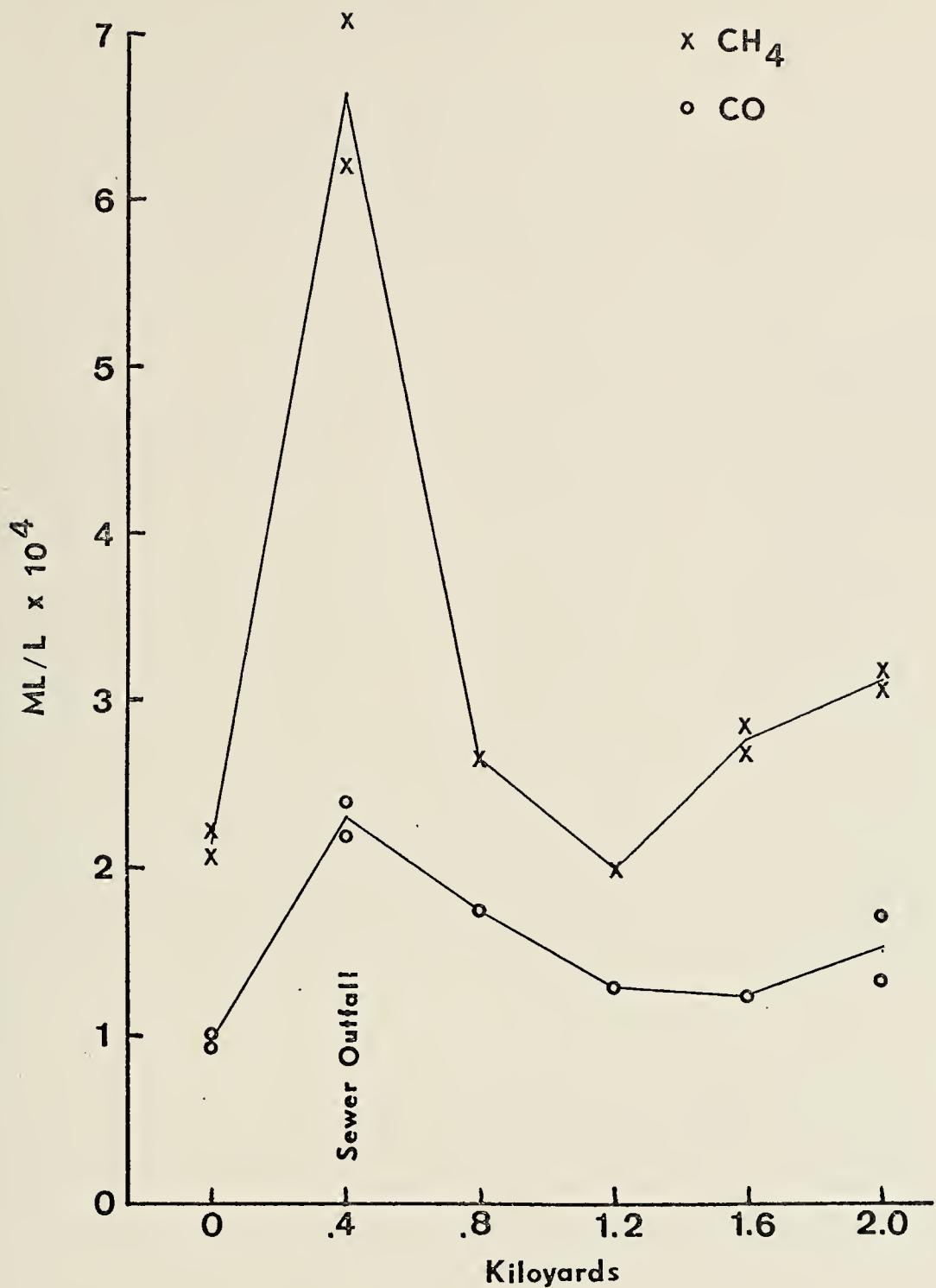


Figure 16. CH_4 and CO concentrations for a transect along Del Monte Beach.

Table 9
 Methane and Carbon Monoxide
 Concentrations for the Transect Along
 Del Monte Beach

Date	Location	Methane (ml/l)	Carbon Monoxide (ml/l)
11-16-72	Kelp Edge	2.214×10^{-4}	0.935×10^{-4}
11-16-72	Kelp Edge	2.089×10^{-4}	1.040×10^{-4}
11-16-72	Sewer Outfall	7.053×10^{-4}	2.380×10^{-4}
11-16-72	Sewer Outfall	6.198×10^{-4}	2.191×10^{-4}
11-16-72	Pump House	2.631×10^{-4}	1.763×10^{-4}
11-16-72	Beach lab	1.995×10^{-4}	1.276×10^{-4}
11-16-72	Apartments	2.823×10^{-4}	1.222×10^{-4}
11-16-72	Apartments	2.673×10^{-4}	1.241×10^{-4}
11-16-72	Public Beach	3.161×10^{-4}	1.172×10^{-4}
11-16-72	Public Beach	3.054×10^{-4}	3.054×10^{-4}

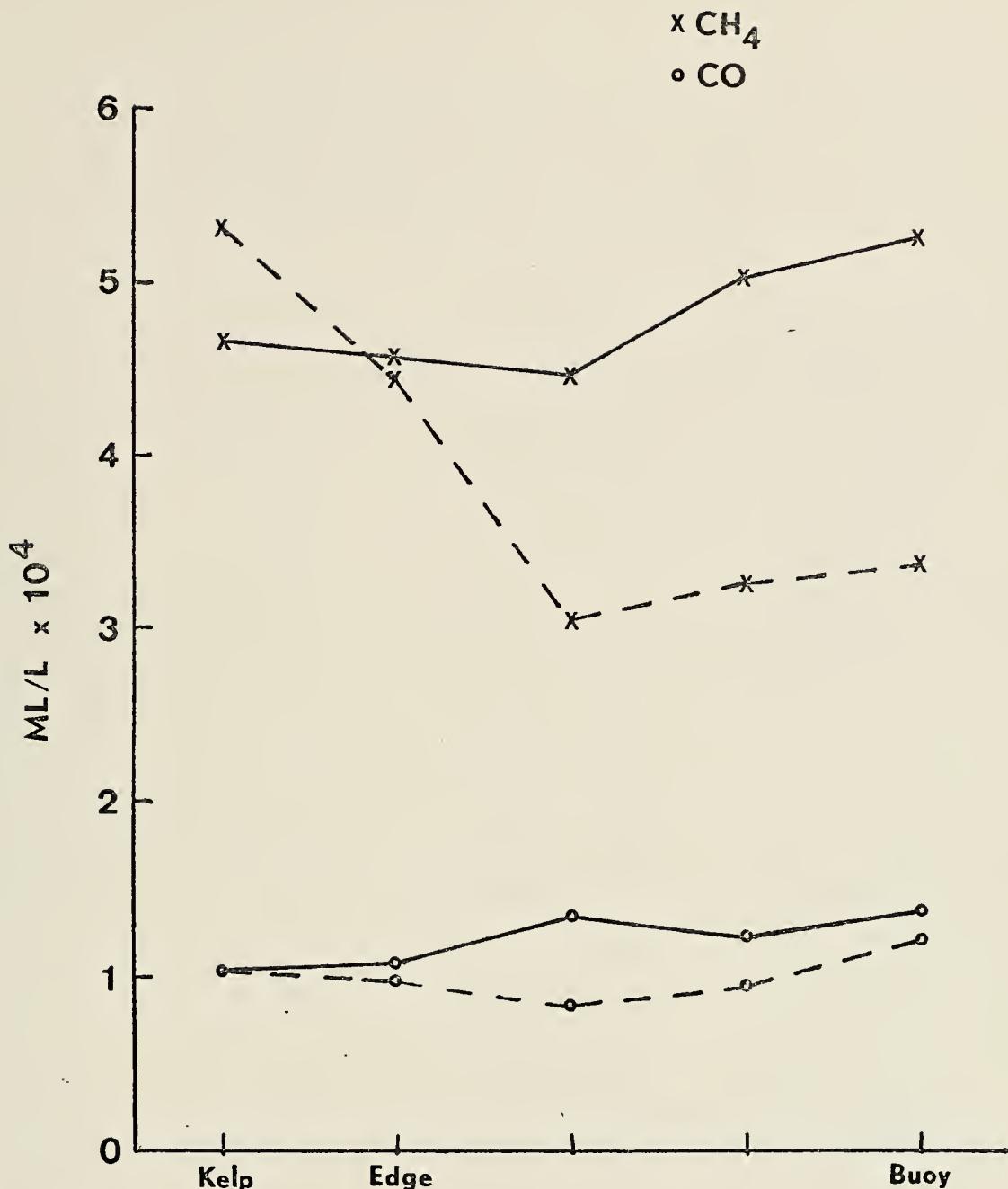


Figure 17. CH_4 and CO concentrations from transects from Del Monte Beach to the R-4 Buoy. Solid line is data for 1 December 1972 and dashed line is data for 14 December 1972.

Table 10
 Methane and Carbon Monoxide
 Concentrations for the Transect from
 Del Monte Beach to the R-4 Buoy

Date	Location	Methane (ml/l)	Carbon Monoxide (ml/l)
12-01-72	Station 1	4.682×10^{-4}	1.039×10^{-4}
12-01-72	Kelp Edge	4.517×10^{-4}	1.041×10^{-4}
12-01-72	Kelp Edge	4.609×10^{-4}	1.080×10^{-4}
12-01-72	Open Water 1	4.490×10^{-4}	1.330×10^{-4}
12-01-72	Open Water 1	4.433×10^{-4}	1.389×10^{-4}
12-01-72	Open Water 2	4.965×10^{-4}	1.272×10^{-4}
12-01-72	Open Water 2	5.105×10^{-4}	1.271×10^{-4}
12-01-72	Station 6	5.287×10^{-4}	1.416×10^{-4}
12-01-72	Station 6	5.233×10^{-4}	1.397×10^{-4}
12-14-72	Station 1	5.358×10^{-4}	1.046×10^{-4}
12-14-72	Station 1	5.277×10^{-4}	1.082×10^{-4}
12-14-72	Kelp Edge	4.374×10^{-4}	0.990×10^{-4}
12-14-72	Kelp Edge	4.594×10^{-4}	1.143×10^{-4}
12-14-72	Open Water 1	2.924×10^{-4}	0.856×10^{-4}
12-14-72	Open Water 1	3.120×10^{-4}	0.854×10^{-4}
12-14-72	Open Water 2	3.277×10^{-4}	0.997×10^{-4}
12-14-72	Station 6	3.382×10^{-4}	1.223×10^{-4}

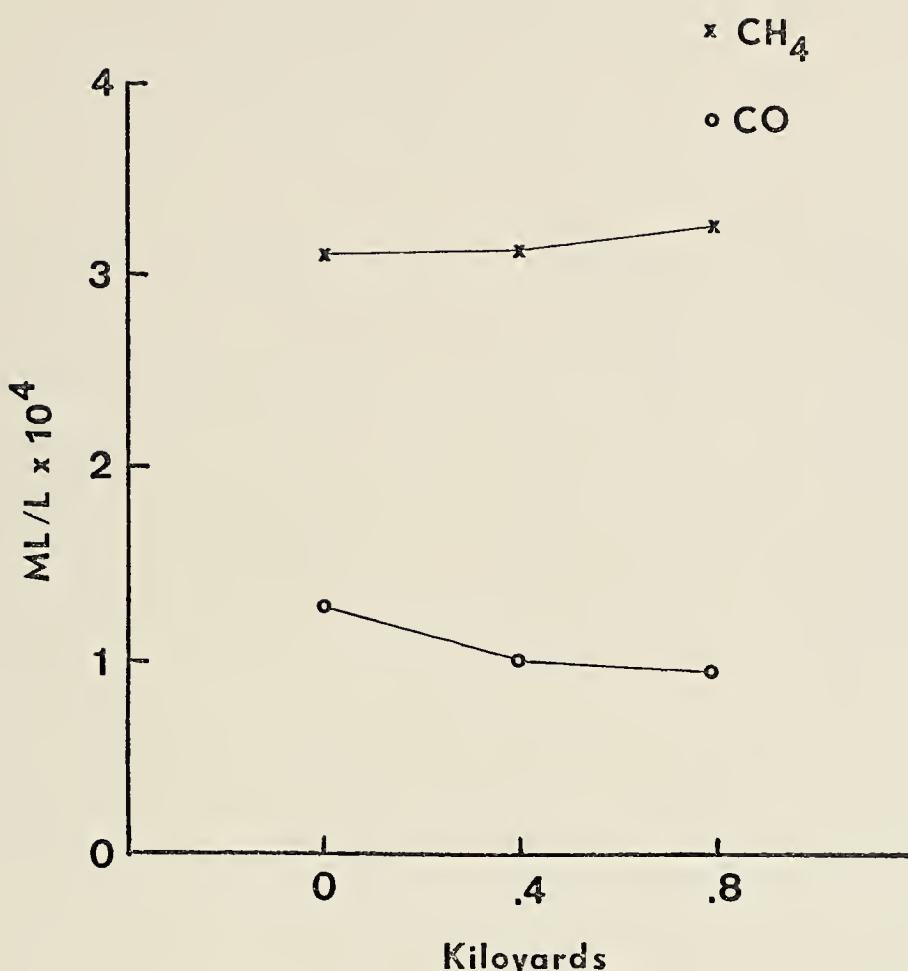


Figure 18. CH_4 and CO concentrations for a transect from Point Cabrillo to the R-4 Bell Buoy.

Table 11
 Methane and Carbon Monoxide
 Concentrations for the Transect from
 Point Cabrillo to the R-4 Buoy

<u>Date</u>	<u>Location</u>	<u>Methane</u> (ml/l)	<u>Carbon Monoxide</u> (ml/l)
12-15-72	Station 2	3.117×10^{-4}	1.242×10^{-4}
12-15-72	Station 2	3.096×10^{-4}	1.305×10^{-4}
12-15-72	Intermediate	3.091×10^{-4}	0.970×10^{-4}
12-15-72	Intermediate	3.136×10^{-4}	1.024×10^{-4}
12-15-72	Station 6	3.231×10^{-4}	0.932×10^{-4}

V. DISCUSSION OF RESULTS

A. PRECISION AND ACCURACY

The precision and accuracy of the gas chromatograph system was checked by running ten consecutive calibration checks. The average of these ten runs was computed as well as the RMS deviation. It was found that to be within plus or minus two RMS deviations, errors as large as 10% could be expected for carbon monoxide and 7.8% for methane.

These errors came mainly from the peak area measurements. First, the determination of a base line proved a problem if there was any drift in the system. To minimize this, the same method of baseline determination was used in each run. This may not eliminate the error, but it should keep it constant.

Area calculation by peak height and width at half height measurement only approximates actual peak area. To solve this problem, either the very careful use of a planimeter, cutting out the peaks and weighing them, or an automatic integrator could be used. These first two unfortunately have the problem of increased analysis time, and the last one, high cost.

B. OPEN OCEAN DEEP STATION

The comparison of Figures 7 and 8 shows a striking correspondence between carbon monoxide concentrations and primary productivity. These results seem to give support to the hypothesis that phytoplankton may be producing carbon monoxide, but this is unexpected and requires experimental verification.

Some anaerobic, methanogenic bacteria are known to convert carbon monoxide to methane (Pine 1971, Jaffe 1970). Figure 19 shows an increase

in both methane and carbon monoxide from the surface to 15 meters. If the carbon monoxide production is related to phytoplankton, and if methane is produced anaerobically in situ (or carried there from an anaerobic source) this could explain the concurrent increase of these gases in the surface layer. From 15-100 meters, the carbon monoxide decreases, but the methane continues to increase. This may show continued bacterial conversion of carbon monoxide to methane below the layer of high productivity, or the advection of a methane rich water mass in the upper 100 meters.

Below 100 meters both gases decrease with depth.

C. TEMPORAL STUDY

The temporal study showed a marked difference between the stations in Monterey Bay (Stations 1 and 2) and those on the exposed coast. Station 1 at Del Monte Beach showed great variability in both methane and carbon monoxide (Figure 9). This was expected since there are many more factors affecting these gases in this environment; sewage disposal plant effluent, pollution from Monterey harbor, and city storm sewer outlets all contribute to the water budget of this station. It is interesting to compare the values of primary productivity with the methane concentrations. They both show marked decrease following the onset of the first storm and then both increase throughout the rest of the study. Carbon monoxide does not show any correlation with primary productivity at this station. The cause for the large jump in carbon monoxide concentration during the period 16-28 November 1972 is unknown. It is possible that this variability in concentrations was present at the other stations and that the more frequent sampling of Del Monte Beach merely showed the transient nature of these gases.

Point Cabrillo (Station 2), still being within the confines of the

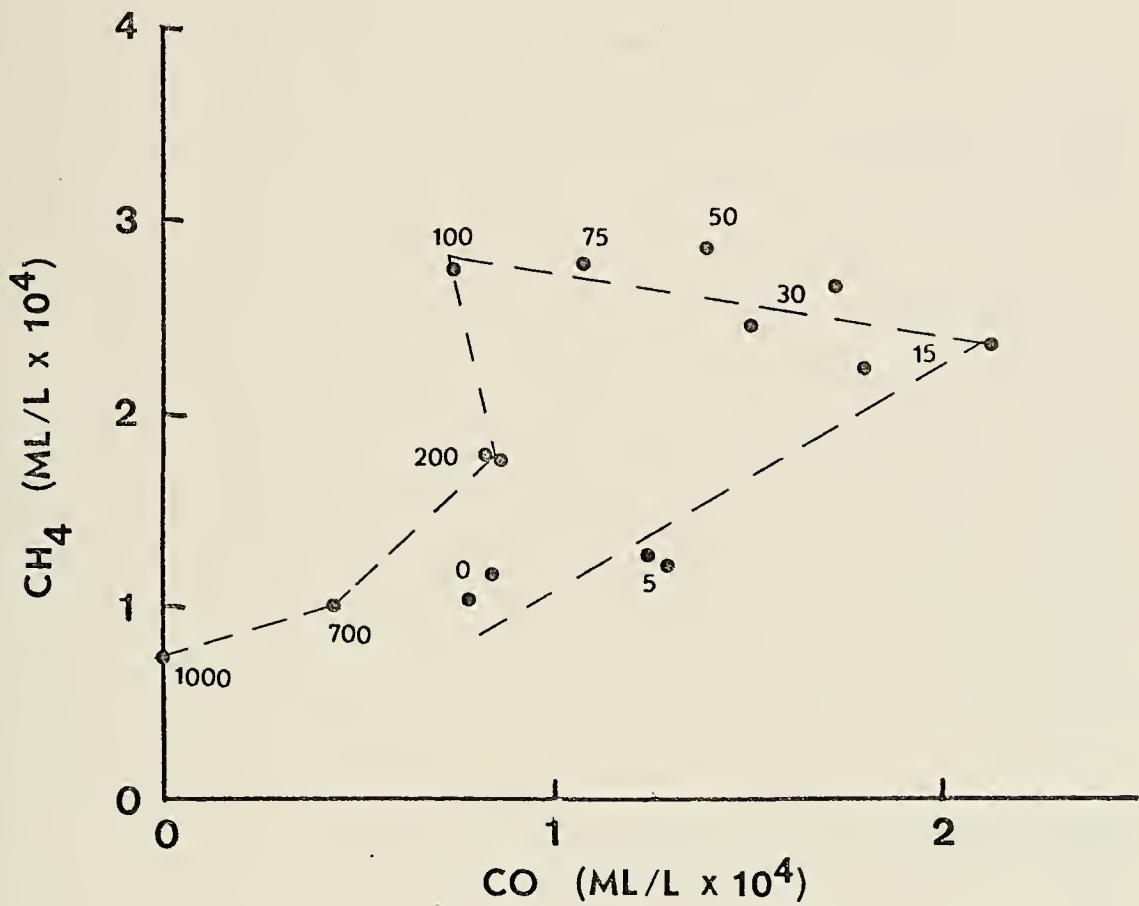


Figure 19. Plot of CH_4 vs. CO in the Deep Ocean Station in the Monterey Canyon. Numbers indicate depth in meters.

bay, had characteristics similar to those found at Del Monte Beach. Since it was further removed from the sources of the pollutants at Del Monte Beach, the variability was not as great. There was still a parallel trend between methane and primary productivity, and none noted for carbon monoxide.

Station 3 at Point Pinos North is subject to the heavy surf action from oceanic waves. Like stations 1 and 2, it, too, showed close correlation between methane and primary productivity, but it also showed a trend between carbon monoxide and primary productivity after the heavy rains.

Station 4 was located just off the Pacific Grove sewage disposal plant outfall at Point Pinos. Here, the methane-primary productivity and the post storm carbon monoxide-primary productivity correlations were both good. It was interesting to note that little or no influence from the sewage effluent was noted in the kelp bed off shore, while the rocks near the outfall appeared to be devoid of life. This probably shows the diluting effect on the heavily chlorinated sewage of the turbulent mixing from the surf in this area.

Station 5 at Point Joe initially showed exceptionally high methane concentrations. Following the first storm, this was reduced to values comparable to the other four stations. Following the rains, it increased more rapidly than any other station. These high methane concentrations seem to indicate a methane source near Point Joe. It is very possible that this source is the Pebble Beach sewage outfall, located just south of Point Joe, where apparently untreated sewage is dumped directly into the ocean.

When the methane concentration was exceptionally high, primary productivity was low. This does not follow the pattern set by the other four

stations. It is possible that phytoplankton growth is inhibited when methane concentrations reach a certain toxic level, or both may be affected by a third factor. After the rains, and subsequent decrease in methane concentrations, the correlation between methane and primary productivity was again apparent.

Carbon monoxide and primary productivity showed very good correlations throughout the sampling period at this station.

D. GRADIENT ANALYSIS

The gradient analysis (Figure 15) indicates the possibility of methane sources at Del Monte Beach and Point Joe. Since the high methane gradients exist at or near sewage disposal plant outfalls, this supports the hypothesis that methane can be used as a tracer for sewage effluent.

The carbon monoxide gradient was not as dramatic as that for methane. It also showed the highest value at Point Joe. An examination of the plant population of Point Joe may help explain this. Stations 1 through 4 are predominantly Macrocystis pyrifera while station 5 is mainly Nereocystis leutkeana. Loewus and Delwiche (1963) have shown that Nereocystis leutkeana produces three times as much carbon monoxide as Macrocystis pyrifera. Therefore, one would expect higher carbon monoxide concentrations at Point Joe than at other stations. A closer examination of the macroalgae at each station might help to explain the variation in the carbon monoxide gradients.

The open ocean value of 1.1×10^{-4} ml/l measured in the Monterey Canyon is 2.4 times the equilibrium concentration of methane in seawater (assuming methane content of air is 1.24 ppm). Similarly, the carbon monoxide concentration of 0.81×10^{-4} ml/l is 22.5 times the equilibrium concentration (assuming the highest reported oceanic partial pressure for

carbon monoxide of 0.17 ppm). The average concentration for both carbon monoxide and methane at each station was considerably higher than the oceanic value. Therefore, it is likely that these kelp beds are sources of these gases to the atmosphere.

E. TRANSECTS

The transects from Del Monte Beach to the R-4 buoy showed the effect of the rains on these gases. Before the rains, the concentrations of both gases in the kelp bed and in the open water were high. After the rains, the values in the kelp were high, but the open water values were decreased substantially. This may show the effect of the kelp as an inhibitor to the mixing process.

The transect from Point Cabrillo to the R-4 buoy showed essentially constant, low methane concentrations. This transect was taken after the rains, and since no methane source near Point Cabrillo was indicated in the gradient analysis, this type of profile could be expected.

The carbon monoxide, on the other hand, showed a decrease seaward. If, indeed, the macroalgae are producing carbon monoxide, this decrease away from the kelp bed could be expected.

Since the gradient analysis indicated a source of methane at Del Monte Beach, a transect was taken there to determine the source of the pollutant (Figure 16). As was expected, extremely high methane values were found in the "boil" above the sewage outfall. Since Monterey does not use an anaerobic treatment process, the source of this methane raised an interesting question.

One possibility was that organic matter was settling out around the outfall. This could create anaerobic conditions and thus generate high concentrations of methane. This idea was abandoned after talking with divers who had frequented the area.

The second premise was that the methane was already present before the sewage reached the treatment plant. This was substantiated by some early attempts to analyze tap water. The city water had so much methane in it that the electrometer in the gas chromatograph was immediately saturated and no value could be obtained.

There was also an indication that the harbor was also a source of both carbon monoxide and methane. With the leakage of petroleum products and the exhaust from power boats, it is no wonder that this was observed.

VI. SUMMARY

Results from the open ocean station and several nearshore stations show a possible correlation between carbon monoxide and primary productivity in this relatively pollution-free environment. Methane was also shown to be present in the upper 100 meters, showing a maximum at about 50 meters. This may indicate biological production of this gas in an oxygen rich environment.

At the nearshore stations, concentrations of methane and carbon monoxide were both far above equilibrium and open ocean values. It was shown that methane concentrations were highest in the vicinity of sewage outfalls and that the source of this gas may be the city water supply. Carbon monoxide concentrations may very well be dependent on the type of plant life in the surrounding waters.

Rain was shown to act as a depressant for methane concentrations and primary productivity, but had little effect on carbon monoxide. Methane in very high concentrations appeared to act as a poison to phytoplankton since primary productivity dropped dramatically in methane rich waters.

The transects showed that methane was an effective tracer for sewage disposal plant effluent and that the kelp seemed to slow the diffusion and mixing of pollutants.

From equilibrium considerations, the waters of Monterey Bay were found to be supersaturated with carbon monoxide and methane. The highest concentrations were, for the most part, located in the kelp beds. These results indicate that, at least locally, the ocean is a source of carbon monoxide and methane in the atmosphere.

VII. RECOMMENDATIONS

Further work with this system could prove valuable in the field of air-sea interaction. By incorporating an air sampling loop, direct measurements of interfacial gradients for a variety of components could be made. It would have been especially helpful in the present study to have known the atmospheric concentrations of these gases.

Due to the limitations of the gas chromatograph used, the hydrocarbons higher than methane that were trapped were not analyzed. In the future, a useful addition to the system would be a dual channel, dual detector gas chromatograph. This would allow simultaneous analysis of a wide variety of organics. The addition of a self-integrating recorder would provide the capability of making on-station, real-time measurements of these dissolved gases.

It is recommended that further work on carbon monoxide and methane be pursued in the Monterey Submarine Canyon. Paucity of ship time aboard R/V Acania precluded further investigation of this area during this study.

APPENDIX A

Analysis Procedures

Sample Transfer

1. Bottle on.
2. Open drain valve to sample bottle to remove old sample from line.
3. Turn drain valve to fill stripping chamber.
4. Turn transfer valve.
5. At 100 ml. turn off transfer valve.
6. Turn drain valve off at zero.
7. Open pressure release and close again.

Stripping

1. Place cold traps.
2. Valve T1 to Trap.
3. Valve T2 to Trap.
4. Trap/Anal valve to Trap.
5. Cal valve to Cal Fill.
6. Trap/Bypass to Trap.
7. Open trap isolation valves.
8. Ensure cap is on valve T1.
9. Open helium purge and start timer.
10. Turn on magnetic stirring bar.
11. Check exhaust flow (13).
12. Close trap isolation valves at 720 sec.
13. Drain 25 - 50 ml. from chamber.

14. Switch Trap/Bypass valve to Bypass.
15. Leave helium flow on.

Analysis

1. Remove cold traps.
2. Place hot water on Trap 2.
3. Warm for one minute.
4. Place Trap/Anal valve to Anal.
5. Set range and attenuation.
6. Zero recorder using Bucking Voltage (obtain baseline).
7. Turn recorder on Low (1"/min.)
8. Turn valve T2 to anal at reference line on recorder.
9. Analysis is complete in seven minutes.

Clean up

1. Drain stripping chamber to about 475 ml.
2. Turn off purge helium.
3. Drain to 500 ml.
4. Place hot water on Trap 1.
5. Remove cap from T1.
6. Switch T1 to Anal.
7. Warm for 1 minute.
8. Repalce cap on T1.
9. Remove hot water.
10. Remove sample bottle.

APPENDIX B

Procedures for Blanks and Calibrations

Calibration (direct injection)

Valve T1 - Trap position.

Valve T2 - Trap position

Anal/Trap - Trap position

Cal valve - Cal Fill position

Trap/Bypass - Bypass position

Fill cal loop

Cal valve - Cal inj. position

Calibration (through traps)

Valve T1 - Trap position

Valve T2 - Trap position

Anal/Trap - Anal position

Cal valve - Cal Fill position

Trap/Bypass - Trap position

Fill Cal Loop

Cal valve - Cal inj. position

Run as if it were a sample, follow stripping and analysis procedure as in Appendix A.

Blank (Run before a series of samples)

Valve T1 - Trap position

Valve T2 - Trap position

Trap/Bypass - Trap position

Cal valve - not important

Run degassed sample, follow stripping and analysis procedures in Appendix A to clear system or check for purge gas contamination.

Appendix C

Program for Computing Concentrations

```

DIMENSION H(2),CH(2),W(2),CW(5),A(2),CA(2),CV(2),VOL(2),CONC(2),
1COUNT(2),CCOUNT(2),LOC(30),Y(30),DEPTH(30)
CV(1)=70.3E-6
CV(2)=71.1E-6
J=1
K=1000
WRITE(6,102)
102 FORMAT(12.0,DATE,12X,'LOCATION',12X,'BOTT. NO.',12X,'DEPTH',12X,
1'METHANE',12X,'CARBON MONOXIDE',/1
1000 READ(5,100,END=9999) IDATE,IBOT,LOC,J,DEPTH,CH(1),CW(1),CRNG,
1CATT,H(1),W(1),RNG,ATT,SAMP
100 FORMAT(16.14,5A4,11,14,8F5.0,F5.3)
A(1)=H(1)*W(1)
CA(1)=CH(1)*CW(1)
COUNT(1)=A(1)*RNG*ATT
CCOUNT(1)=CA(1)*CRNG*CATT
VOL(1)=(CV(1)*COUNT(1))/CCOUNT(1)
CONC(1)=VOL(1)/SAMP
1F(1-EQ-1) GO TO 1000
101 WRITE(6,101) IDATE,IBOT,LOC,J,DEPTH,CONC(1),CONC(2),
1FORMAT(10.14,T11.16,T22,5A4,T14,T169,14,T150,14,T185,1PE10.3,
101 GO TO 1000
9999 STOP
END

```


DISSOLVED GASES DATA SHEET

Raw Data^a

DISSOLVED GASES DATA SHEET

R. V. Anderson 9

CRUISE
VESSEL

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U. S. NAVAL POSTGRADUATE SCHOOL
DEPARTMENT OF OCEANOGRAPHY

DISSOLVED GASES DATA SHEET

DISSOLVED GASES DATA SHEET

K. V. Acciaia		CRUISE NO.		DATE/TIME COLLECTED		DATE/TIME ANALYZED		ANALYST			
DATE	BOTT. NO.	LOCATION		CAL	H CAL	W CAL	RANGE CAL	ATTEN CAL	RANGE COMP	ATTEN COMP	VOLUME (L)
11/23/62	21	10	11	12	13	14	15	16	17	18	19
11/23/62	22	20	21	22	23	24	25	26	27	28	29
11/23/62	23	30	31	32	33	34	35	36	37	38	39
11/23/62	24	40	41	42	43	44	45	46	47	48	49
11/23/62	25	50	51	52	53	54	55	56	57	58	59
11/23/62	26	60	61	62	63	64	65	66	67	68	69
11/23/62	27	70	71	72	73	74	75	76	77	78	79
11/23/62	28	80	81	82	83	84	85	86	87	88	89
11/23/62	29	90	91	92	93	94	95	96	97	98	99
11/23/62	30	100	101	102	103	104	105	106	107	108	109
11/23/62	31	110	111	112	113	114	115	116	117	118	119
11/23/62	32	120	121	122	123	124	125	126	127	128	129
11/23/62	33	130	131	132	133	134	135	136	137	138	139
11/23/62	34	140	141	142	143	144	145	146	147	148	149
11/23/62	35	150	151	152	153	154	155	156	157	158	159
11/23/62	36	160	161	162	163	164	165	166	167	168	169
11/23/62	37	170	171	172	173	174	175	176	177	178	179
11/23/62	38	180	181	182	183	184	185	186	187	188	189
11/23/62	39	190	191	192	193	194	195	196	197	198	199
11/23/62	40	200	201	202	203	204	205	206	207	208	209
11/23/62	41	210	211	212	213	214	215	216	217	218	219
11/23/62	42	220	221	222	223	224	225	226	227	228	229
11/23/62	43	230	231	232	233	234	235	236	237	238	239
11/23/62	44	240	241	242	243	244	245	246	247	248	249
11/23/62	45	250	251	252	253	254	255	256	257	258	259
11/23/62	46	260	261	262	263	264	265	266	267	268	269
11/23/62	47	270	271	272	273	274	275	276	277	278	279
11/23/62	48	280	281	282	283	284	285	286	287	288	289
11/23/62	49	290	291	292	293	294	295	296	297	298	299
11/23/62	50	300	301	302	303	304	305	306	307	308	309
11/23/62	51	310	311	312	313	314	315	316	317	318	319
11/23/62	52	320	321	322	323	324	325	326	327	328	329
11/23/62	53	330	331	332	333	334	335	336	337	338	339
11/23/62	54	340	341	342	343	344	345	346	347	348	349
11/23/62	55	350	351	352	353	354	355	356	357	358	359
11/23/62	56	360	361	362	363	364	365	366	367	368	369
11/23/62	57	370	371	372	373	374	375	376	377	378	379
11/23/62	58	380	381	382	383	384	385	386	387	388	389
11/23/62	59	390	391	392	393	394	395	396	397	398	399
11/23/62	60	400	401	402	403	404	405	406	407	408	409
11/23/62	61	410	411	412	413	414	415	416	417	418	419
11/23/62	62	420	421	422	423	424	425	426	427	428	429
11/23/62	63	430	431	432	433	434	435	436	437	438	439
11/23/62	64	440	441	442	443	444	445	446	447	448	449
11/23/62	65	450	451	452	453	454	455	456	457	458	459
11/23/62	66	460	461	462	463	464	465	466	467	468	469
11/23/62	67	470	471	472	473	474	475	476	477	478	479
11/23/62	68	480	481	482	483	484	485	486	487	488	489
11/23/62	69	490	491	492	493	494	495	496	497	498	499
11/23/62	70	500	501	502	503	504	505	506	507	508	509
11/23/62	71	510	511	512	513	514	515	516	517	518	519
11/23/62	72	520	521	522	523	524	525	526	527	528	529
11/23/62	73	530	531	532	533	534	535	536	537	538	539
11/23/62	74	540	541	542	543	544	545	546	547	548	549
11/23/62	75	550	551	552	553	554	555	556	557	558	559
11/23/62	76	560	561	562	563	564	565	566	567	568	569
11/23/62	77	570	571	572	573	574	575	576	577	578	579
11/23/62	78	580	581	582	583	584	585	586	587	588	589
11/23/62	79	590	591	592	593	594	595	596	597	598	599
11/23/62	80	600	601	602	603	604	605	606	607	608	609
11/23/62	81	610	611	612	613	614	615	616	617	618	619
11/23/62	82	620	621	622	623	624	625	626	627	628	629
11/23/62	83	630	631	632	633	634	635	636	637	638	639
11/23/62	84	640	641	642	643	644	645	646	647	648	649
11/23/62	85	650	651	652	653	654	655	656	657	658	659
11/23/62	86	660	661	662	663	664	665	666	667	668	669
11/23/62	87	670	671	672	673	674	675	676	677	678	679
11/23/62	88	680	681	682	683	684	685	686	687	688	689
11/23/62	89	690	691	692	693	694	695	696	697	698	699
11/23/62	90	700	701	702	703	704	705	706	707	708	709
11/23/62	91	710	711	712	713	714	715	716	717	718	719
11/23/62	92	720	721	722	723	724	725	726	727	728	729
11/23/62	93	730	731	732	733	734	735	736	737	738	739
11/23/62	94	740	741	742	743	744	745	746	747	748	749
11/23/62	95	750	751	752	753	754	755	756	757	758	759
11/23/62	96	760	761	762	763	764	765	766	767	768	769
11/23/62	97	770	771	772	773	774	775	776	777	778	779
11/23/62	98	780	781	782	783	784	785	786	787	788	789
11/23/62	99	790	791	792	793	794	795	796	797	798	799
11/23/62	100	800	801	802	803	804	805	806	807	808	809
11/23/62	101	810	811	812	813	814	815	816	817	818	819
11/23/62	102	820	821	822	823	824	825	826	827	828	829
11/23/62	103	830	831	832	833	834	835	836	837	838	839
11/23/62	104	840	841	842	843	844	845	846	847	848	849
11/23/62	105	850	851	852	853	854	855	856	857	858	859
11/23/62	106	860	861	862	863	864	865	866	867	868	869
11/23/62	107	870	871	872	873	874	875	876	877	878	879
11/23/62	108	880	881	882	883	884	885	886	887	888	889
11/23/62	109	890	891	892	893	894	895	896	897	898	899
11/23/62	110	900	901	902	903	904	905	906	907	908	909
11/23/62	111	910	911	912	913	914	915	916	917	918	919
11/23/62	112	920	921	922	923	924	925	926	927	928	929
11/23/62	113	930	931	932	933	934	935	936	937	938	939
11/23/62	114	940	941	942	943	944	945	946	947	948	949
11/23/62	115	950	951	952	953	954	955	956	957	958	959
11/23/62	116	960	961	962	963	964	965	966	967	968	969
11/23/62	117	970	971	972	973	974	975	976	977	978	979
11/23/62	118	980	981	982	983	984	985	986	987	988	989
11/23/62	119	990	991	992	993	994	995	996	997	998	999
11/23/62	120	1000	1001	1002	1003	1004	1005	1006	1007	1008	1009
11/23/62	121	1010	1011	1012	1013	1014	1015	1016	1017	1018	1019
11/23/62	122	1020	1021	1022	1023	1024	1025	1026	1027	1028	1029
11/23/62	123	1030	1031	1032	1033	1034	1035	1036	1037	1038	1039
11/23/62	124	1040	1041	1042	1043	1044	1045	1046	1047	1048	1049
11/23/62	125	1050	1051	1052	1053	1054	1055	1056	1057	1058	1059
11/23/62	126	1060	1061	1062	1063	1064	1065	1066	1067	1068	1069
11/23/62	127	1070	1071	1072	1073	1					

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110' Kent 12 160000 Min 72 161400 Net 72

ANALYST DATE/TIME ANALYZED
NO. DATE/TIME ANALYZED

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40' Boat	13	17000 Nov 72	171400 Nov 72	ANALYST
VESSEL	CRUISE NO.	DATE/TIME COLLECTED	DATE/TIME	ANALYZED

VESSEL	DATE	CRUISE NO.	LOCATION	DATE/TIME	COLLECTED	DATE/TIME	ANALYZED	ANALYST
BOTT. NO.	COM	DEPTH (A)	COM	H CAL	W CAL	RANGE CAL	H COMP	RANGE COMP

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ANALYST _____ DATE/TIME _____ ANALYZED _____

DISSOLVED GASES DATA SHEET

110' *Scout*

20

CRUISE NO.

Analyst

Fahey

140830 Dec 72

141200 Dec 72

DATE/TIME ANALYZED

Analyst

VESSEL	BOTT. NO.	LOCATION		DATE/TIME COLLECTED	DATE/TIME ANALYZED	ANALYST									
		CO ₂ DEPTH (M)	H ₂ CAL				W ₂ CAL	RANGE CAL	ATTEN CAL	H ₂ COMP	W ₂ COMP	RANGE COMP	ATTEN COMP	VOLUME (L)	
121472	1	CALCOFIL A		151516171819202122232425262728292031	151516171819202122232425262728292031		151516171819202122232425262728292031	151516171819202122232425262728292031	151516171819202122232425262728292031	151516171819202122232425262728292031	151516171819202122232425262728292031	151516171819202122232425262728292031	151516171819202122232425262728292031	151516171819202122232425262728292031	151516171819202122232425262728292031
	2														
	3	CALCOFIL (Quartz)													
	4														
	5														
	6														
	7														
	8	CALCOFIL (NOMEX)													
	9														
	10														
	11														
	12														
	13														
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and Other Nearshore Habitats

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ABSTRACT

A system for the determination of dissolved gases in seawater by gas chromatography was constructed and used to find the concentrations of methane and carbon monoxide in a variety of habitats around the Monterey Peninsula. Methane was shown to have a maximum of 2.8×10^{-4} ml/l at 50 meters at the open ocean station, with a surface value of 1.1×10^{-4} ml/l. The surface waters at the nearshore stations were almost three times this value. Methane was also shown to be an effective tracer for sewage effluent. The carbon monoxide maximum of 2.1×10^{-4} ml/l was found at 15 meters which correlated closely with primary productivity (Rowley 1973). The surface values of 0.81×10^{-4} ml/l was lower than the nearshore values. All stations sampled were found to be highly supersaturated with both gases. This indicates that in this area, the ocean is a major source of both methane and carbon monoxide.

KEY WORDS	LINK A		LINK B		LINK C	
	ROLE	WT	ROLE	WT	ROLE	WT
gas chromatography						
methane						
carbon monoxide						
dissolved gases						
gradient analysis						
gas analysis						
seawater analysis						
pollution						

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